

| AQA Chemistry (8462) from 2016 Topics C4.1 Atomic structure and the periodic table              |                                                                                                                                                                                                                   |   |   |   |
|-------------------------------------------------------------------------------------------------|-------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------|---|---|---|
| Topic                                                                                           | Student Checklist                                                                                                                                                                                                 | R | A | G |
| 4.1.1 A simple model of the atom, symbols, relative atomic mass, electronic charge and isotopes | State that everything is made of atoms and recall what they are                                                                                                                                                   |   |   |   |
|                                                                                                 | Describe what elements and compounds are                                                                                                                                                                          |   |   |   |
|                                                                                                 | State that elements and compounds are represented by symbols; and use chemical symbols and formulae to represent elements and compounds                                                                           |   |   |   |
|                                                                                                 | Write word equations and balanced symbol equations for chemical reactions, including using appropriate state symbols                                                                                              |   |   |   |
|                                                                                                 | <b>HT ONLY:</b> Write balanced half equations and ionic equations                                                                                                                                                 |   |   |   |
|                                                                                                 | Describe what a mixture is                                                                                                                                                                                        |   |   |   |
|                                                                                                 | Name and describe the physical processes used to separate mixtures and suggest suitable separation techniques                                                                                                     |   |   |   |
|                                                                                                 | Describe how the atomic model has changed over time due to new experimental evidence, inc discovery of the atom and scattering experiments (inc the work of James Chadwick)                                       |   |   |   |
|                                                                                                 | Describe the difference between the plum pudding model of the atom and the nuclear model of the atom                                                                                                              |   |   |   |
|                                                                                                 | State the relative charge of protons, neutrons and electrons and describe the overall charge of an atom                                                                                                           |   |   |   |
|                                                                                                 | State the relative masses of protons, neutrons and electrons and describe the distribution of mass in an atom                                                                                                     |   |   |   |
|                                                                                                 | Calculate the number of protons, neutrons and electrons in an atom when given its atomic number and mass number                                                                                                   |   |   |   |
|                                                                                                 | Describe isotopes as atoms of the same element with different numbers of neutrons                                                                                                                                 |   |   |   |
|                                                                                                 | Define the term relative atomic mass and why it takes into account the abundance of isotopes of the element                                                                                                       |   |   |   |
|                                                                                                 | Calculate the relative atomic mass of an element given the percentage abundance of its isotopes                                                                                                                   |   |   |   |
| 4.1.2 The periodic table                                                                        | Describe how electrons fill energy levels in atoms, and represent the electron structure of elements using diagrams and numbers                                                                                   |   |   |   |
|                                                                                                 | Recall how the elements in the periodic table are arranged                                                                                                                                                        |   |   |   |
|                                                                                                 | Describe how elements with similar properties are placed in the periodic table                                                                                                                                    |   |   |   |
|                                                                                                 | Explain why elements in the same group have similar properties and how to use the periodic table to predict the reactivity of elements                                                                            |   |   |   |
|                                                                                                 | Describe the early attempts to classify elements                                                                                                                                                                  |   |   |   |
|                                                                                                 | Explain the creation and attributes of Mendeleev's periodic table                                                                                                                                                 |   |   |   |
|                                                                                                 | Identify metals and non-metals on the periodic table, compare and contrast their properties                                                                                                                       |   |   |   |
|                                                                                                 | Explain how the atomic structure of metals and non-metals relates to their position in the periodic table                                                                                                         |   |   |   |
|                                                                                                 | Describe noble gases (group 0) and explain their lack of reactivity                                                                                                                                               |   |   |   |
|                                                                                                 | Describe the properties of noble gases, including boiling points, predict trends down the group and describe how their properties depend on the outer shell of electrons                                          |   |   |   |
|                                                                                                 | Describe the reactivity and properties of group 1 alkali metals with reference to their electron arrangement and predict their reactions                                                                          |   |   |   |
|                                                                                                 | Describe the properties of group 7 halogens and how their properties relate to their electron arrangement, including trends in molecular mass, melting and boiling points and reactivity                          |   |   |   |
|                                                                                                 | Describe the reactions of group 7 halogens with metals and non-metals                                                                                                                                             |   |   |   |
|                                                                                                 | <b>Chem ONLY:</b> Describe the properties of transition metals and compare them with group 1 elements, including melting points and densities, strength and hardness, and reactivity (for Cr, Mn Fe, Co, Ni & Cu) |   |   |   |

# Pixl Club 1mao Paper 1

**Siddappa N.Byrareddy**



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