

Assessment

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Understanding Concepts

11. What is a stable electron configuration?
12. What does each dot in an electron dot diagram represent?
13. What process changes atoms into ions?
14. What keeps the ions in their fixed positions within a crystal lattice?
15. What are subscripts used for in chemical formulas?
16. Explain why a melted ionic compound is a good conductor of electric current, but a solid ionic compound is a poor conductor of electric current.
17. What distinguishes single, double, and triple covalent bonds?
18. Explain why the covalent bonds in molecules of O_2 and N_2 are always nonpolar.
19. Explain why, in a covalent bond between oxygen and hydrogen, the hydrogen atom has a partial positive charge and the oxygen atom has a partial negative charge.
20. What is the name of the binary compound formed from potassium and iodine?
21. Write the formulas for the compounds called copper(I) chloride and copper(II) chloride.
22. Name the compounds represented by the space-filling models labeled A, B, and C.



23. In general, what determines the strength of metallic bonds?
24. What properties of copper and tin change when these metals are mixed together to form bronze?
25. What advantage of magnesium is retained in magnesium alloys? What disadvantage is reduced?

Chemical Bonds 187

Assessment



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Reviewing Content

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Understanding Concepts

11. In a stable electron configuration, the highest occupied energy level is filled with electrons.
12. Each dot represents a valence electron.
13. The transfer of electrons.
14. Attractions between neighboring cations and anions keep the ions in fixed positions within the lattice.
15. A subscript is used to show the number of atoms of an element in a molecule or the ratio of ions in a crystal lattice.
16. When an ionic compound melts, ions can move away from their fixed locations in the crystal lattice.
17. Two atoms share two electrons in a single bond, four in a double bond, and six in a triple bond.
18. The covalent bonds in molecules of elements are always nonpolar because the atoms have the same attraction for electrons.
19. The oxygen atom has a greater attraction for electrons than the hydrogen atom does.
20. Potassium iodide
21. CuCl and CuCl_2
22. A is sulfur trioxide, B is carbon monoxide, and C is nitrogen dioxide
23. In general, the more valence electrons a metal has, the stronger the metal bonds are.
24. A mixture of copper and tin is harder and stronger than either metal in its pure form.
25. The advantage that is retained is that magnesium is a lightweight metal. The disadvantage that is reduced is that magnesium is a soft metal.

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R Barnett



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