



Lab Assistant

# Classifying Chemical Reactions

Directions: Write the type of chemical reaction in the space provided next to each equation. Name the product or the products of the reaction at the end.

## Column I

1. \_\_\_\_\_ 1. A precipitate forms in a gas liquid collection over water (combustion, decomposition, precipitation)
2. \_\_\_\_\_ 2. Two reactants exchange products and/or anions/cations
3. \_\_\_\_\_ 3. A hydrocarbon explodes, usually in a combustion
4. \_\_\_\_\_ 4. Chlorine atoms break down into many pieces (decomposition)
5. \_\_\_\_\_ 5. A type of synthesis reaction that produces one molecule

Directions: Write the chemical formula for each of the reactants below in the space provided.

combustion \_\_\_\_\_ 6. ethane + oxygen (ethane)

decomposition \_\_\_\_\_ 7.  $\text{H}_2\text{O} \rightarrow \text{H}_2 + \text{O}_2$  (hydrogen + oxygen)

synthesis

synthesis reaction

\_\_\_\_\_ 8.  $\text{Mg} + \text{O}_2 \rightarrow \text{MgO}$  (magnesium + oxygen)

single displacement reaction \_\_\_\_\_ 9.  $\text{Fe} + \text{CuSO}_4 \rightarrow \text{FeSO}_4 + \text{Cu}$  (iron + copper sulfate)

synthesis reaction

\_\_\_\_\_ 10.  $\text{H}_2 + \text{O}_2 \rightarrow \text{H}_2\text{O}$  (hydrogen + oxygen)

\_\_\_\_\_ 11.  $\text{C}_2\text{H}_6 + \text{O}_2 \rightarrow \text{CO}_2 + \text{H}_2\text{O}$  (ethane + oxygen)

decomposition

single displacement reaction

\_\_\_\_\_ 12.  $\text{H}_2\text{SO}_4 + \text{NaOH} \rightarrow \text{Na}_2\text{SO}_4 + \text{H}_2\text{O}$  (sulfuric acid + sodium hydroxide)

synthesis reaction

single displacement reaction

\_\_\_\_\_ 13.  $\text{H}_2 + \text{Cl}_2 \rightarrow \text{HCl}$  (hydrogen + chlorine)

decomposition

single displacement reaction

\_\_\_\_\_ 14.  $\text{C}_2\text{H}_6 + \text{O}_2 \rightarrow \text{CO}_2 + \text{H}_2\text{O}$  (ethane + oxygen)

single displacement reaction \_\_\_\_\_ 15.  $\text{H}_2 + \text{O}_2 \rightarrow \text{H}_2\text{O}$  (hydrogen + oxygen)

single displacement reaction

## Section 2 Reinforcement Classifying Chemical Reactions

**RJ Alexander**



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