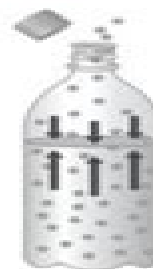


15. Describe the two diagrams of a bottled carbonated beverage below as *greater pressure* or *lower pressure*, and then as *greater solubility* or *lower solubility*. How do these two examples illustrate the relationship between the solubility of a gas and its vapor pressure?

In the closed bottle, the pressure is high and lots of CO_2 is dissolved in the
carbonated beverage. Once the cap is removed, the vapor pressure decreases,
allowing more CO_2 in the carbonated beverage to escape.



greater pressure and greater solubility lower pressure and lower solubility

16. How does a solution become supersaturated? If the temperature of a solution having
a small excess of solid solute is raised, the solute dissolves. If the solution is then
allowed to cool slowly, the excess solute may stay dissolved at a temperature below
the temperature at which it would ordinarily crystallize.

SECTION 16.2 CONCENTRATIONS OF SOLUTIONS (pages 480–486)

This section explains how to solve problems involving molarity of a solution, how to prepare dilute solutions from more concentrated solutions, and what is meant by percent by volume and percent by mass.

► Molarity (pages 480–482)

1. A measure of the amount of solute dissolved in a given quantity of solvent is the concentration of a solution.
2. The most important unit of concentration in chemistry is molarity.
3. Is the following sentence true or false? Molarity is the number of moles of dissolved solute per liter of solvent. false

Section 16 3 Colligative Properties Of Solutions Key

Annelies Wilder-Smith



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