

Practice Problems

Do your best to solve the following problems.

SECTION 12.1 THE MOLE: A MEASUREMENT OF MATTER

- Calculate the molar mass of each of the following compounds:
 - phosphorus pentachloride (PCl_5)
 - carbon tetrachloride (CCl_4)
- Calculate the mass percent of each of the following compounds:
 - KNO_3
 - Na_2SO_4
- Calculate the number of moles of each of the following compounds:
 - 10.0 g KNO_3
 - 10.0 g Na_2SO_4
- Calculate the mass of each of the following compounds:
 - 0.500 mol KNO_3
 - 0.500 mol Na_2SO_4
- Calculate the mass of 1.00 mol of each of the following compounds:
 - KNO_3
 - Na_2SO_4

SECTION 12.2 MOLE—MASS AND MOLE—VOLUME RELATIONSHIPS

- Calculate the molar mass of each of the following compounds:

a. H_2O	b. CO_2	c. H_2SO_4	d. KNO_3
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- Calculate the mass in grams of each of the following:

a. 0.500 mol H_2O	b. 0.500 mol CO_2
c. 0.500 mol H_2SO_4	d. 0.500 mol KNO_3
e. 0.500 mol Na_2SO_4	f. 0.500 mol CaCl_2
- Calculate the mass in grams of each of the following compounds:
 - 0.500 mol H_2O
 - 0.500 mol CO_2
 - 0.500 mol H_2SO_4
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 - 0.500 mol KNO_3
 - 0.500 mol Na_2SO_4
 - 0.500 mol CaCl_2

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