

Relative Strengths of **Acids** and **Bases** in Aqueous Solutions at 25°C

Name of Acid	Acid		Base	K_a
Hydriodic	HI	→	$H^+ + I^-$	$>> 1$
Hydrobromic	HBr	→	$H^+ + Br^-$	$>> 1$
Hydrochloric	HCl	→	$H^+ + Cl^-$	$>> 1$
Sulfuric	H_2SO_4	→	$H^+ + HSO_4^-$	> 1
Nitric	HNO_3	→	$H^+ + NO_3^-$	> 1
Hydronium Ion	H_3O^+	⇌	$H^+ + H_2O$	5.50×10^{-1}
Oxalic	$H_2C_2O_4$	⇌	$H^+ + HC_2O_4^-$	5.90×10^{-2}
Sulfurous	H_2SO_3	⇌	$H^+ + HSO_3^-$	1.41×10^{-2}
Hydrogen Sulfate Ion	HSO_4^-	⇌	$H^+ + SO_4^{2-}$	1.20×10^{-2}
Phosphoric	H_3PO_4	⇌	$H^+ + H_2PO_4^-$	7.52×10^{-3}
Nitrous	HNO_2	⇌	$H^+ + NO_2^-$	5.62×10^{-4}
Hydrofluoric	HF	⇌	$H^+ + F^-$	3.53×10^{-4}
Formic	HCO_2H	⇌	$H^+ + HCO_2^-$	1.78×10^{-4}
Benzoic	$C_6H_5CO_2H$	⇌	$H^+ + C_6H_5CO_2^-$	6.46×10^{-5}
Acetic	CH_3CO_2H	⇌	$H^+ + CH_3CO_2^-$	1.76×10^{-5}
Carbonic	H_2CO_3	⇌	$H^+ + HCO_3^-$	4.30×10^{-7}
Hydrogen Sulfite Ion	HSO_3^-	⇌	$H^+ + SO_3^{2-}$	1.02×10^{-7}
Dihydrogen Phosphate Ion	$H_2PO_4^-$	⇌	$H^+ + HPO_4^{2-}$	6.23×10^{-8}
Boric	H_3BO_3	⇌	$H^+ + H_2BO_3^-$	5.79×10^{-10}
Ammonium Ion	NH_4^+	⇌	$H^+ + NH_3$	5.64×10^{-10}
Hydrogen Carbonate Ion	HCO_3^-	⇌	$H^+ + CO_3^{2-}$	5.61×10^{-11}
Monohydrogen Phosphate Ion	HPO_4^{2-}	⇌	$H^+ + PO_4^{3-}$	2.20×10^{-13}
Water	H_2O	⇌	$H^+ + OH^-$	1.00×10^{-14}

Strong Acids

Weak Acids

Weakest Acids

Weakest Bases

Weak Bases

Stronger Bases

All reactions assume water as a reactant.



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$$K_a = [H^+][A^-] / [HA]$$

$$pK_a = -\log K_a$$

$$pH = -\log [H^+]$$

Section 2 Study Guid Strength Of Acid And Bases

Mark Steyvers



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