

## 19.5

## HYDROGEN IONS AND ACIDITY

## Section Review

## Objectives

1. Calculate concentrations based on the pH of acids and bases and vice versa.
2. Calculate pH and pOH and concentrations of acids and bases and vice versa.
3. Calculate the pH of weak acids.

## Vocabulary

- |                        |                             |                   |
|------------------------|-----------------------------|-------------------|
| 1. autoionization      | 2. conjugate acid-base pair | 3. weak acid/base |
| 4. ionization constant | 5. acid rain                | 6. neutralization |
|                        | 7. acid solution            | 8. pH             |

## Key Equation

- |                                     |                                     |
|-------------------------------------|-------------------------------------|
| 1. $\text{pH} = -\log [\text{H}^+]$ | 2. $\text{pH} = 14 - \text{pOH}$    |
| 3. $\text{pH} = -\log [\text{H}^+]$ | 4. $\text{pH} = -\log [\text{H}^+]$ |

## Part A Completion

Write the requested unit or numerical information of the values provided that are indicated in parentheses. Significant figures are indicated with the appropriate symbol.

1. What substance is \_\_\_\_\_, a weak hydrogen ion donor?
2. \_\_\_\_\_ is the concentration of ions which participate in ionic equilibrium. \_\_\_\_\_ is \_\_\_\_\_.
3. \_\_\_\_\_ is the \_\_\_\_\_ of \_\_\_\_\_.
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20. \_\_\_\_\_ is the \_\_\_\_\_ of \_\_\_\_\_.

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# Section 19 2 Hydrogen Ions And Acidity Practice Problems

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