

## 19.5

## HYDROGEN IONS AND ACIDITY

## Section Review

## Objectives

1. Calculate concentrations based on the pH of acids and bases and vice versa.
2. Calculate pH and pOH and concentrations of acids and bases and vice versa.
3. Calculate the pH of weak acids and bases.

## Vocabulary

- |                   |                             |                       |
|-------------------|-----------------------------|-----------------------|
| 1. autoionization | 2. conjugate acid-base pair | 3. acid solution      |
| 4. autoionization | 5. acid rain                | 6. acid rain solution |
|                   | 7. acid solution            | 8. pH                 |

## Key Equation

- |                                     |                                     |
|-------------------------------------|-------------------------------------|
| 1. $\text{pH} = -\log [\text{H}^+]$ | 2. $\text{pH} = -\log [\text{H}^+]$ |
| 3. $\text{pH} = -\log [\text{H}^+]$ | 4. $\text{pH} = -\log [\text{H}^+]$ |

## Part A Completion

Write the equation and/or the chemical formula(s) of the species involved in the reaction described in each problem. Be sure to include the correct state symbols.

1. Write the equation for the autoionization of water.
2. Write the equation for the autoionization of water.
3. Write the equation for the autoionization of water.
4. Write the equation for the autoionization of water.
5. Write the equation for the autoionization of water.
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18. Write the equation for the autoionization of water.
19. Write the equation for the autoionization of water.
20. Write the equation for the autoionization of water.

1.  $\text{H}_2\text{O} \rightleftharpoons \text{H}^+ + \text{OH}^-$ 2.  $\text{H}_2\text{O} \rightleftharpoons \text{H}^+ + \text{OH}^-$ 3.  $\text{H}_2\text{O} \rightleftharpoons \text{H}^+ + \text{OH}^-$ 4.  $\text{H}_2\text{O} \rightleftharpoons \text{H}^+ + \text{OH}^-$ 5.  $\text{H}_2\text{O} \rightleftharpoons \text{H}^+ + \text{OH}^-$ 6.  $\text{H}_2\text{O} \rightleftharpoons \text{H}^+ + \text{OH}^-$ 7.  $\text{H}_2\text{O} \rightleftharpoons \text{H}^+ + \text{OH}^-$ 8.  $\text{H}_2\text{O} \rightleftharpoons \text{H}^+ + \text{OH}^-$ 9.  $\text{H}_2\text{O} \rightleftharpoons \text{H}^+ + \text{OH}^-$ 10.  $\text{H}_2\text{O} \rightleftharpoons \text{H}^+ + \text{OH}^-$ 11.  $\text{H}_2\text{O} \rightleftharpoons \text{H}^+ + \text{OH}^-$ 12.  $\text{H}_2\text{O} \rightleftharpoons \text{H}^+ + \text{OH}^-$ 13.  $\text{H}_2\text{O} \rightleftharpoons \text{H}^+ + \text{OH}^-$ 14.  $\text{H}_2\text{O} \rightleftharpoons \text{H}^+ + \text{OH}^-$ 15.  $\text{H}_2\text{O} \rightleftharpoons \text{H}^+ + \text{OH}^-$ 16.  $\text{H}_2\text{O} \rightleftharpoons \text{H}^+ + \text{OH}^-$ 17.  $\text{H}_2\text{O} \rightleftharpoons \text{H}^+ + \text{OH}^-$ 18.  $\text{H}_2\text{O} \rightleftharpoons \text{H}^+ + \text{OH}^-$ 19.  $\text{H}_2\text{O} \rightleftharpoons \text{H}^+ + \text{OH}^-$ 20.  $\text{H}_2\text{O} \rightleftharpoons \text{H}^+ + \text{OH}^-$

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## **Table of Contents Section 19 2 Hydrogen Ions And Acidity Answer Key**

1. Understanding the eBook Section 19 2 Hydrogen Ions And Acidity Answer Key
  - The Rise of Digital Reading Section 19 2 Hydrogen Ions And Acidity Answer Key
  - Advantages of eBooks Over Traditional Books
2. Identifying Section 19 2 Hydrogen Ions And Acidity Answer Key
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Section 19 2 Hydrogen Ions And Acidity Answer Key
  - User-Friendly Interface
4. Exploring eBook Recommendations from Section 19 2 Hydrogen Ions And Acidity Answer Key
  - Personalized Recommendations
  - Section 19 2 Hydrogen Ions And Acidity Answer Key User Reviews and Ratings
  - Section 19 2 Hydrogen Ions And Acidity Answer Key and Bestseller Lists

5. Accessing Section 19 2 Hydrogen Ions And Acidity Answer Key Free and Paid eBooks
  - Section 19 2 Hydrogen Ions And Acidity Answer Key Public Domain eBooks
  - Section 19 2 Hydrogen Ions And Acidity Answer Key eBook Subscription Services
  - Section 19 2 Hydrogen Ions And Acidity Answer Key Budget-Friendly Options
6. Navigating Section 19 2 Hydrogen Ions And Acidity Answer Key eBook Formats
  - ePub, PDF, MOBI, and More
  - Section 19 2 Hydrogen Ions And Acidity Answer Key Compatibility with Devices
  - Section 19 2 Hydrogen Ions And Acidity Answer Key Enhanced eBook Features
7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Section 19 2 Hydrogen Ions And Acidity Answer Key
  - Highlighting and Note-Taking Section 19 2 Hydrogen Ions And Acidity Answer Key
  - Interactive Elements Section 19 2 Hydrogen Ions And Acidity Answer Key
8. Staying Engaged with Section 19 2 Hydrogen Ions And Acidity Answer Key
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Section 19 2 Hydrogen Ions And Acidity Answer Key
9. Balancing eBooks and Physical Books Section 19 2 Hydrogen Ions And Acidity Answer Key
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Section 19 2 Hydrogen Ions And Acidity Answer Key
10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
11. Cultivating a Reading Routine Section 19 2 Hydrogen Ions And Acidity Answer Key
  - Setting Reading Goals Section 19 2 Hydrogen Ions And Acidity Answer Key
  - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Section 19 2 Hydrogen Ions And Acidity Answer Key
  - Fact-Checking eBook Content of Section 19 2 Hydrogen Ions And Acidity Answer Key
  - Distinguishing Credible Sources
13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
- Exploring Educational eBooks

### 14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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