This is meant as a guide. It does NOT contain every single thing you need to know for the final exam. Review notes, benchmark review sheets, homeworks, chapter study guides, etc.

Bring your book to school by Monday for turn in (seniors should bring book to bookstore by themselves ASAP)

Gas Laws (Ch. 13):

Subtopics: atmospheric pressure, Dalton's law of partial pressures, combined gas law, Ideal gas law, gas stoichiometry (22.4 L/mol for a gas at STP)

Practice questions:

 If a gas occupies 3.8 L at a pressure of 2.71 atm, what would the pressure be if the volume changes to 1.47 L? (Temp. and amt. of gas constant)

V1=3.8 L; P1=2.71 atm V2=1.47L; P2=? Atm Since all units are standard (L, atm), we do not need to convert the units before proceeding. (P1V1)/T1=(P2V2)/T2 T goes away (not mentioned in problem) P1V1=P2V2 Isolate for P2: P2= (P1V1)/V2 P2= (2.71 atm x 3.8 L)/(1.47 L)= 7.0 atm (L cancelled and it's 2 sig figs)

A gas has a volume of 1.49 L at a temperature of 34.75 degrees Celsius. What would the volume be at 78.41 degrees Celsius? (pressure & amt. of gas constant)

V1= 1.49 L, T1= 34.75 degrees Celsius V2= ? L, T2= 78.41 degrees Celsius We need to convert from Celsius to Kelvins before proceeding. T1= 34.75 + 273= 307.75 K T2= 78.41 + 273= 351.41 K (P1V1)/T1= (P2V2)/T2 Get rid of P V1/T1= V2/T2 Isolate V2: V2= (V1/T1)T2 V2= (1.49 L/307.75 K)(351.41 K)= 1.70 L

 What volume is occupied by 8.47 g of hydrogen gas at 84.7 degrees Celsius and 1.04 atm?

The combined gas law won't work here. You only have one condition and nothing changes.

PV=nRT will work. (n= moles and R= 0.08026 L atm/K mol)

P= 1.04 atm, V= 7 L, n= OOPS we need to convert from grams to moles!, T= 84.7 degrees Celcius + 273= 357.7 K

n: Hydrogen gas is diatomic!

8.47 g H₂ x (1 mol H₂/2.02 g H₂)= 4.193069307 mol H₂

Isolate V: V= (nRT)/P= (4.193069307 mol H₂ x 0.08206 L atm/K mol x 357.7 K)/ 1.04 atm= 118 L

4. What volume is occupied by 56.75 g of oxygen gas at STP?

Second Semester Standard Chemistry Review Guide

David Kirk

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