Chemistry 2nd Semester Exam Review Sheet: Chapters 9, 10, 11, 12, 13, 14, 15, and 16

Chapter 9: Chemical Quantities

- Solve stoichiometry problems using grams →moles → moles → grams calculations. Examples below:
 - Calculate moles/grams of (NH₄):SO₄ produced when 30.0 moles of NH₃ react with H₂SO₄.

2NH3 + H2SO4 → (NH4)2SO4 30.0 mol × moles

30.0 moles x 1mol/2mol = 15.0 moles

15.0 moles (NH4)2SO4 x 132.2 grams/mole = 1983 = 1980 grams (to 3 s.f.)

b. Na₂CO₃ + 2HNO₂ → 2NaNO₃ + CO₂ + H₂O. How many moles/grams of sodium nitrate are produced by the complete reaction of 100, grams of sodium carbonate? Moles of nitric acid needed to produce 500.0 grams of sodium nitrate? Be able to calculate grams or moles of any reactant or product given a quantity of reactant or product.

Grams → Moles → Moles → Grams

100. g Na2CO3 x 1 mol/106.0 g x 2 mol NaNO3 / 1 mol Na2CO3 x 85.00 g/mol = 160.4 g

500.0 g NaNO3 x 1 mol/85.00 g x 2 mol / 2 mol x 63.02 g/mol = 370.7 grams

Calculate percent yield of products. Example below using (b) above.

What is the theoretical yield of NaNO₃ if 100.0 grams of Na₂CO₃ are used in the reaction? What is the percent yield of NaNO₃ if 130.5 grams of NaNO₃ are actually collected?

Theoretical yield = 160.4 grams (answer from above)

Percent Yield = (Actual Yield/Theoretical Yield) x 100 = (130.5 g/160.4 g) x 100 = 81.36%

What is a limiting reactant in a chemical reaction? The reactant that is used up first, thereby limiting the amount of product that can be formed.

Chapter 10: Modern Atomic Theory

4. Describe Rutherford's model of the atom. What could it not explain?

Nucleus (very small) consisting of protons with electrons in orbits around the nucleus. Could not explain how the electrons remained in orbit.

- Define electromagnetic radiation, wavelength, and frequency.
- How do atoms gain energy to become excited? Absorb a photon of energy.

.How do they lose energy to return to their ground state? Release a photon of same energy as that absorbed.

 What is the evidence that energy absorbed by the hydrogen atom is quantized? Specific wavelengths of light seen in emission spectra for atoms. (page 285-86)

Describe the Bohr model of the atom. How did it improve on Rutherford's model?

Same as Rutherford with respect to nucleus. Electrons located on orbits at specific distances from nucleus.

9. Draw the orbital diagram (with the arrows) for an element: for example, phosphorus, sodium, and fluorine.

See book, notes, probs 54, 96

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Scott C. Dulebohn

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