



Seat Cordoba 2005 Manual

SEAT Servicio de Asistencia Técnica

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web the balanced chemical equation for this reaction is $\text{Mg}(\text{s}) + 2\text{HNO}_3(\text{aq}) \rightarrow \text{Mg}(\text{NO}_3)_2(\text{aq}) + \text{H}_2(\text{g})$ in this equation the s symbolizes solid magnesium aq represents aqueous nitric acid aq signifies magnesium nitrate in a dissolved state and g represents the gaseous hydrogen produced

$\text{HNO}_3 + \text{Mg} + \text{C}_2\text{H}_3\text{O}_2 \rightarrow \text{H}_2\text{C}_2\text{H}_3\text{O}_2 + \text{Mg}(\text{NO}_3)_2$ chemical equation - Aug 11 2023

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web mar 4 2018 $\text{Mg} + 2\text{HNO}_3 \rightarrow \text{Mg}(\text{NO}_3)_2 + \text{H}_2$ on the left side there are Mg 1 magnesium in elemental form H_2 there are 2 molecules of nitric acid each acid has one hydrogen so two molecules of nitric acid is 2 hydrogens N_2 there are 2 molecules of nitric acid each acid has one nitrogen so two molecules of nitric acid is 2 nitrogens O_6

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web magnesium is already balanced on the reactant and product side balancing nitrogen by multiplying HNO_3 by 2 we get $\text{Mg} + 2\text{HNO}_3 \rightarrow \text{Mg}(\text{NO}_3)_2 + \text{H}_2$ magnesium 2 nitric acid $\text{Mg}(\text{NO}_3)_2$ s magnesium nitrate H_2 g hydrogen gas oxygen and hydrogen get automatically balanced by balancing nitrogen

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