

CHEMISTRY 1211 TEST #3

Use a Scantron sheet for Multiple Choices (3pts each). Do not write on the Periodic Table.

Use back of this page to show your work/answer for the Big Question.

1mi=1.609km; 1lbs=453.6g; 1gal=3.785L; 1mL=1cm³; 1=10⁻¹⁰⁰ m;
1cal=4.184J; g=9.807m/s
1atm=760mmHg; Density(Hg)=13.60g/mL; c=2.9979x10⁸m/s; h=6.626x10⁻³⁴J R_∞=2.180x10⁻¹⁸J;
m_e=9.109x10⁻³¹ kg; R=8.315J/k.mol.

1. Which photon has the lowest energy ?

A(X-ray), B(Ultraviolet radiation, UV), C(Visible light), D(Infrared radiation, IR), **E(Microwave)**.

2. Energy of an electromagnetic wave increases with (a) increasing, (b) decreasing, frequency; therefore a blue light has a higher energy than a red light because of its (c) higher, (d) lower frequency than that of a red light. **A(a,c)**, B(a,d), C(b,c), D(b,d).

3. Frequency of a radio wave is 104.7 MHz. It's wavelength must be about A(3 fm), **B(3 cm)**, C(3 m), D(3 km), E(3 Mm).

4. Find energy of an X-ray with a wavelength of 1.0x10⁻¹⁰ m (= 0.10 nm). Report it (a) per photon and (b) per mole.

A(1.986x10⁻¹²/photon; 1.196x10⁶kJ/mol), B(1.986x10⁻¹²J/photon; 1.196x10⁶kJ/mol), C(1.986x10⁻¹²J/photon; 1.196x10⁶kJ/mol), D(1.986x10⁻¹²J/photon; 1.196x10⁶kJ/mol), **E(1.986x10⁻¹²J/photon; 1.196x10⁶kJ/mol)**.

5. Who got the Nobel Prize in 1921 with an idea that light behave as if it is a particle whose energy is directly proportional to its frequency ?

A(Balmer), B(Planck), **C(Einstein)**, D(Bohr), E(Heisenberg).

6. What is the energy (J) of an electron at the third energy level (i.e., in the L shell with the principle quantum number n=3) in a H atom?

Problem Set Chemistry Test Bank

Zhe Xu



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Table of Contents Problem Set Chemistry Test Bank

1. Understanding the eBook Problem Set Chemistry Test Bank
 - The Rise of Digital Reading Problem Set Chemistry Test Bank
 - Advantages of eBooks Over Traditional Books
2. Identifying Problem Set Chemistry Test Bank
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Problem Set Chemistry Test Bank
 - User-Friendly Interface
4. Exploring eBook Recommendations from Problem Set Chemistry Test Bank
 - Personalized Recommendations
 - Problem Set Chemistry Test Bank User Reviews and Ratings
 - Problem Set Chemistry Test Bank and Bestseller Lists

5. Accessing Problem Set Chemistry Test Bank Free and Paid eBooks
 - Problem Set Chemistry Test Bank Public Domain eBooks
 - Problem Set Chemistry Test Bank eBook Subscription Services
 - Problem Set Chemistry Test Bank Budget-Friendly Options
6. Navigating Problem Set Chemistry Test Bank eBook Formats
 - ePub, PDF, MOBI, and More
 - Problem Set Chemistry Test Bank Compatibility with Devices
 - Problem Set Chemistry Test Bank Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Problem Set Chemistry Test Bank
 - Highlighting and Note-Taking Problem Set Chemistry Test Bank
 - Interactive Elements Problem Set Chemistry Test Bank
8. Staying Engaged with Problem Set Chemistry Test Bank
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Problem Set Chemistry Test Bank
9. Balancing eBooks and Physical Books Problem Set Chemistry Test Bank
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Problem Set Chemistry Test Bank
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Problem Set Chemistry Test Bank
 - Setting Reading Goals Problem Set Chemistry Test Bank
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Problem Set Chemistry Test Bank
 - Fact-Checking eBook Content of Problem Set Chemistry Test Bank
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
- Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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