

Stoichiometry Worksheet

I. ONE REACTANT IN EXCESS – COMBUSTION

You burn an excess of butane (C₄H₁₀) in a limited amount of oxygen (O₂). If 67.2 L of CO₂ gas is formed, what volume of oxygen gas was present at the start of the reaction?

STEP 1: Balance the reaction:



STEP 2: Use a molar conversion to convert quantity of product given to moles:

$$67.2 \text{ L CO}_2 * \frac{(1 \text{ mol CO}_2)}{(22.4 \text{ L CO}_2)} = 3 \text{ mol CO}_2$$

STEP 3: Use a molar ratio to calculate the moles of reactant that I'm asking about:

$$3 \text{ mol CO}_2 * \frac{(13 \text{ mol O}_2)}{(8 \text{ mol CO}_2)} = 4.875 \text{ mol O}_2$$

STEP 4: Use a molar conversion to convert moles of reactant to the units I ask for:

$$4.875 \text{ mol O}_2 \frac{(22.4 \text{ L O}_2)}{(1 \text{ mol O}_2)} = 109.2 \text{ L O}_2 = \boxed{109 \text{ L O}_2}$$

II. ONE REACTANT IN EXCESS – SINGLE DISPLACEMENT

14 g of lithium is combined with excess K₂SO₄. What is the mass of the new ionic compound that is formed? How do you know that this reaction will take place?

STEP 1: Balance the reaction:



Reaction occurs because Li is higher up than K on the activity series.

STEP 2: Use a molar conversion to convert quantity of reactant given to moles:

$$14 \text{ g Li} * \frac{(1 \text{ mol Li})}{(7 \text{ g Li})} = 2 \text{ mol Li}$$

STEP 3: Use a molar ratio to calculate moles of product that I'm asking about:

$$2 \text{ mol Li} * \frac{(1 \text{ mol Li}_2\text{SO}_4)}{(2 \text{ mol Li})} = 1 \text{ mol Li}_2\text{SO}_4$$

STEP 4: Use a molar conversion to convert moles of product to units that I ask for:

$$1 \text{ mol Li}_2\text{SO}_4 * \frac{(110 \text{ g Li}_2\text{SO}_4)}{(1 \text{ mol Li}_2\text{SO}_4)} = \boxed{110 \text{ Li}_2\text{SO}_4}$$

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