

Practice Problems

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SECTION 12.1 THE MOLE: A MEASUREMENT OF MATTER

1. Calculate the molar mass of each of the following compounds:

a. Sodium chloride (table salt or halite) NaCl (molar mass)

b. phosphorus pentachloride (PCl₅) (molar mass)

c. carbon tetrachloride (CCl₄) (molar mass)

2. Calculate the mass percent of each element in the compound:

a. NaCl (molar mass)

b. PCl₅ (molar mass)

3. Determine the molar mass of each compound of the form M_2X_3 :

7. Manganese (molar mass 54.94 g/mol) and sulfur (molar mass 32.07 g/mol)

8. Nitrogen (molar mass 14.01 g/mol) and oxygen (molar mass 16.00 g/mol)

SECTION 12.2 MOLE—MASS AND MOLE—VOLUME RELATIONSHIPS

1. Calculate the mass of each sample of the following compounds:

a. 1.00 mol NaCl (molar mass)

b. 0.500 mol PCl₅ (molar mass)

c. 0.250 mol CCl₄ (molar mass)

d. 0.00100 mol NaCl (molar mass)

2. Calculate the mass of each sample of each of the following:

a. 0.500 mol sodium chloride (NaCl) (molar mass)

b. 0.00100 mol sodium chloride (NaCl) (molar mass)

c. 0.00100 mol sodium chloride (NaCl) (molar mass)

d. 0.00100 mol sodium chloride (NaCl) (molar mass)

e. 0.00100 mol sodium chloride (NaCl) (molar mass)

3. Determine the mass of each sample of the following compounds (molar mass):

a. 0.00100 mol sodium chloride (NaCl) (molar mass)

4. Determine the mass of each sample of each of the following:

a. 0.00100 mol sodium chloride (NaCl) (molar mass)

b. 0.00100 mol sodium chloride (NaCl) (molar mass)

c. 0.00100 mol sodium chloride (NaCl) (molar mass)

5. Calculate the mass of each sample of the following compounds (molar mass):

a. 0.00100 mol sodium chloride (NaCl) (molar mass)

Section 10 Chemical Quantities Practice Problems

Answers

JE Gale



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