

1. In the list below, circle the properties of acids and box the properties of bases.

|                   |                                  |                         |
|-------------------|----------------------------------|-------------------------|
| Sour taste        | Feel slippery                    | Contains hydroxide ions |
| Turns litmus blue | Reacts with metals to form $H_2$ | Turns litmus red        |
| Bitter taste      | Contains hydrogen ions           | Corrosive               |

2. Complete the chart below:

| Acid-Base Definitions |                                  |                                   |
|-----------------------|----------------------------------|-----------------------------------|
| Type                  | Acid Definition                  | Base Definition                   |
| Arrhenius             | Donates an $H^+$ ion in solution | Donates an $OH^-$ ion in solution |
| Bronsted-Lowry        | Donates an $H^+$ ion             | Accepts an $H^+$ ion              |
| Lewis                 | Accepts an electron pair         | Donates an electron pair          |

3. What term(s) describe a substance that can act as an acid or a base? **Amphoteric or amphiprotic**
4. Neutralization reactions always produce a **SALT** and **WATER**
5. If the  $[H^+]$  is greater than the  $[OH^-]$ , the solution is acidic basic neutral
6. If the  $[OH^-]$  is greater than the  $[H^+]$ , the solution is acidic basic neutral
7. If the  $[H^+]$  is equal to the  $[OH^-]$ , the solution is acidic basic neutral
8. If the pH is 7, the solution is acidic basic neutral
9. If the pH is greater than 7, the solution is acidic basic neutral
10. If the pH is less than 7, the solution is acidic basic neutral
11. How many moles of lithium hydroxide are required to neutralize 19.2 grams of chloric acid?



$$(19.2 \text{ g } HClO_3) \times (1 \text{ mol } HClO_3/84 \text{ g } HClO_3) \times (1 \text{ mol } LiOH/1 \text{ mol } HClO_3) = \underline{0.228 \text{ mol } LiOH}$$

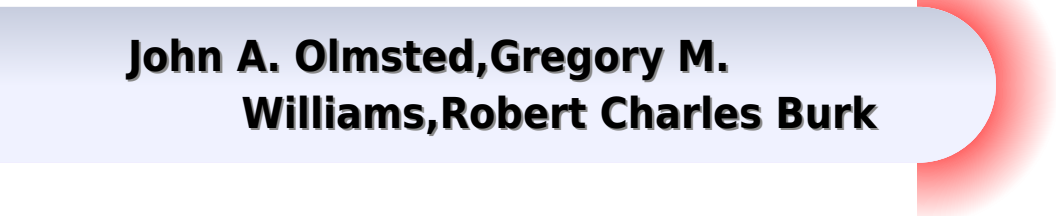
12. How many grams of calcium hydroxide are required to neutralize 132.4 grams of nitric acid?



$$(132.4 \text{ g } HNO_3) \times (1 \text{ mol } HNO_3/63 \text{ g } HNO_3) \times (1 \text{ mol } Ca(OH)_2/2 \text{ mol } HNO_3) \times (74 \text{ g } Ca(OH)_2/1 \text{ mol } Ca(OH)_2) = \underline{77.76 \text{ g } Ca(OH)_2}$$

## Section 20 Acids And Bases Answers

**John A. Olmsted, Gregory M.  
Williams, Robert Charles Burk**



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