An acid is defined as a substance that donates a proton (written H\*) while a base is the substance that receives a proton. Typically the chemical formula can be used to determine the acid, because it will begin with the symbol H. For example in the following equation HCl is the acid and it donates a proton to water.

Acids donate protons Bases accept protons

A proton is a hydrogen ion

In this reaction the HCl is the acid, while the  $H_2O$  acts as the base. This creates two new products: **hydronium**,  $H_3O^*$ , and the chloride ion,  $C\Gamma$ .

Some acids have the ability to donate two or three protons and these are known as **diprotic** or **triprotic acids** respectively. For these acids each successive step of hydrogen donation is represented with its own equation. Consider the diprotic acid called carbonic acid, H<sub>2</sub>CO<sub>3</sub>.

First step:	H <sub>2</sub> CO <sub>3</sub>	+ H <sub>2</sub> O -	→ H <sub>3</sub> O*	+	HCO <sub>3</sub>
Second step:	HCO <sub>3</sub>	+ H <sub>2</sub> O -	→ H <sub>3</sub> O*	+	$CO_3{}^{2-}$

-					
=	See Section	STA.	ACCRECATE VALUE OF THE PARTY.	Section 1	le:

Write the chemical equation that shows what happens when HF (acid) is added to water.

- write the equation

HF + H<sub>2</sub>O -> F' + H<sub>2</sub>O'

- check to make sure the atoms and the charge are balanced

Rewrite each equation and label the acid and the base in each reaction.

1.  $NH_3 + H_2O \rightarrow NH_4^* + OH^-$ 

4.  $H_2O + CN^- \rightarrow HCN + OH^-$ 

HI + OH → Γ + H<sub>2</sub>O

OH<sup>-</sup> + NH<sub>1</sub>\* → H<sub>2</sub>O + NH<sub>3</sub>

3.  $HCO_3^- + HNO_3 \rightarrow H_2CO_3 + NO_3^-$ 

6.  $H_2SO_4 + PO_4^{3-} \rightarrow HPO_4^{2-} + HSO_4^{-}$ 

Fill in the following table.

	Acid	Base	Equation
7	HNO <sub>3</sub>	OH-	$HNO_3 + OH^- \rightarrow H_2O + NO_3^-$
8	Consideration III		$CH_3NH_2 + H_2O \rightarrow OH^- + CH_3NH_3$
9	HCN		$HCN + H_2O \rightarrow H_3O^+ + CN^-$
10	HBr	H <sub>2</sub> O	
1.1	HPO <sub>4</sub> <sup>2</sup>	NH,	
12	noncementary III	100 (000)	$OH^- + H_2S \rightarrow H_2O + HS^-$
13	$H_2C_2O_4$	OH.	
14	HC1O	NH <sub>3</sub>	
15	HSO.	$CO_{2}$	

# **Section 19 4 Acid Base Reactions Worksheet Answers**

Thirumalaisamy P. Velavancorresponding

Section 19 4 Acid Base Reactions Worksheet Ans
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