

## 14.3

## HYDROGEN IONS AND ACIDITY

## Section Review

## Objectives

1. Calculate concentrations based on the pH of a solution and the pH of a solution based on concentration.
2. Calculate pH and concentration of acids and bases and the pH of a solution based on concentration.
3. Calculate the pH of a solution (25°C).

## Vocabulary

- |                   |                                             |                     |
|-------------------|---------------------------------------------|---------------------|
| 1. autoionization | 2. hydronium ion ( $\text{H}_3\text{O}^+$ ) | 3. acid solution    |
| 4. autoionization | 5. acid rain                                | 6. neutral solution |
|                   | 7. acid solution                            | 8. pH               |

## Key Equation

- |                                     |                                               |
|-------------------------------------|-----------------------------------------------|
| 1. $\text{pH} = -\log [\text{H}^+]$ | 2. $\text{pH} = -\log [\text{H}_3\text{O}^+]$ |
| 3. $\text{pH} = -\log [\text{H}^+]$ | 4. $\text{pH} = -\log [\text{H}_3\text{O}^+]$ |

## Part A Completion

Write the equation for the autoionization of water, and explain the relationship between the concentration of hydronium ions and hydroxide ions in pure water.

What substance is  $\text{H}^+$  in water? Hydrogen ion is  $\text{H}^+$  in water. The concentration of  $\text{H}^+$  in water is  $1.0 \times 10^{-7}$  M.

What is the pH of a solution that is  $1.0 \times 10^{-7}$  M  $\text{H}^+$ ? The pH of a solution that is  $1.0 \times 10^{-7}$  M  $\text{H}^+$  is 7.0.

What is the pH of a solution that is  $1.0 \times 10^{-7}$  M  $\text{H}^+$ ? The pH of a solution that is  $1.0 \times 10^{-7}$  M  $\text{H}^+$  is 7.0.

1.  $\text{H}_2\text{O} \rightleftharpoons \text{H}^+ + \text{OH}^-$ 2.  $[\text{H}^+] = 1.0 \times 10^{-7}$  M3.  $\text{pH} = 7.0$ 4.  $[\text{H}^+] = 1.0 \times 10^{-7}$  M5.  $\text{pH} = 7.0$ 6.  $1.0 \times 10^{-7}$  M7.  $1.0 \times 10^{-7}$  M8.  $1.0 \times 10^{-7}$  M9.  $1.0 \times 10^{-7}$  M10.  $1.0 \times 10^{-7}$  M11.  $1.0 \times 10^{-7}$  M12.  $1.0 \times 10^{-7}$  M

## Section 19 2 Hydrogen And Ions Acidity Answer Key

**Xiang Xie**



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