

# RATES OF REACTION

## Section A: Summary Notes

### Energy Changes during Chemical Reactions

- Most reactions do not begin until an amount of energy (activation energy) has been added to the reaction mixture.
- The **activation energy** is often called the '**energy hill**' which must be 'overcome' by the addition of this amount of energy before a reaction can take place.
- When activation energy is added to the reactants, a so-called **activated complex** is formed.
- The *activated complex is temporary, unstable, high-energy composition of atoms, which represents a transition state between reactants and the products.*
- When the activated complex is formed during a reaction, this complex can lead either to the formation of new bonds, i.e. molecules of the products, or to re-formation of the old bonds, thereby returning to being reactants.
- For the reaction:



The formation of an activated complex as a transitional step can be represented as follows:



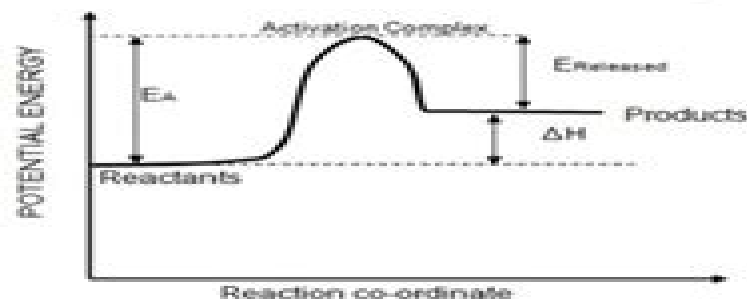
- When an activated complex is formed during a reaction, this complex can lead either to the formation of new bonds, i.e. molecules of the products or to the re-formation of the old bonds, thereby returning to being the reactants. This is **reversibility for the reaction**.

### Heat of Reaction (Enthalpy)

- Heat of the reaction ( $\Delta H$ ) is the difference between the energy of the products and the energy of the reactants.

$$\Delta H = E_{\text{products}} - E_{\text{reactants}}$$

- For an **endothermic reaction**,  $E_{\text{products}} > E_{\text{reactants}}$ , therefore  $\Delta H$  is **positive**.



**Note:** In a reversible reaction, the energy released in forming the products in the forward reaction is the same as the activation energy ( $E_A$ ) of the reverse reaction.

- For an **exothermic reaction**,  $E_{\text{products}} < E_{\text{reactants}}$ ,  $\Delta H$  is **negative**.

## Section 18 Rates Of Reaction 467 Answers

**Raffaela Di Napoli**

A red circular graphic with a gradient, appearing as a semi-circle or a partial circle, located to the right of the author's name.

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