

Quiz & Worksheet - Colligative Properties and Raoult's Law

1. What happens to vapor pressure when you add a solute to a solution?

- ☐ It lowers the vapor pressure.
- ☐ It has no effect.
- ☐ It raises the vapor pressure.
- ☐ It causes the reaction to reach equilibrium.

2. What is osmotic pressure?

- ☐ It is the minimum pressure that must be applied to a solution to stop osmosis from happening.
- ☐ It is the maximum pressure that must be applied to a solution to stop osmosis from happening.
- ☐ It is the maximum pressure of the vapor over a liquid at equilibrium.
- ☐ It is the minimum pressure of the vapor over a liquid at equilibrium.

3.

9.0g of urea (a nonelectrolyte) is dissolved in 360 g water to create a solution. The temperature of the solution is 25.0°C. If the vapor pressure of water at 25.0°C is 23.80 mmHg, calculate the vapor pressure of the urea solution at 25.0°C in mmHg.

The molar mass of urea 60g/mol and water is 18g/mol.

- ☐ 0.99
- ☐ 23.91
- ☐ 23.61
- ☐ 3.57

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Section 18 3 Colligative Properties Of Solutions Answers

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