

## Section 12.3 Limiting Reagent & Percent Yield

### Answer Key

Michelle Ganian

#### Vocabulary

1. e
2. c
3. d
4. a
5. b

1. Answer: 127g Cu  
Because...

Find the limiting reagent:

$$54\text{g Al} \times 1 \text{ mole Al}/27\text{g Al} = 2 \text{ mole Al}$$

$$319\text{g CuSO}_4 \times 1 \text{ mole CuSO}_4/159.5\text{g CuSO}_4 = 2 \text{ mole CuSO}_4 \text{ (limiting reagent)}$$

$$2 \text{ mole CuSO}_4 \times 3 \text{ mole Cu}/3 \text{ mole CuSO}_4 \times 63.5\text{g Cu}/1 \text{ mole Cu} = 127\text{g Cu}$$

(Multiply the founded limiting reagent with moles of Copper (3) over moles of CuSO<sub>4</sub> (3) from the equation to cancel out moles of CuSO<sub>4</sub>.)

Then, to cancel out 3 moles Cu, multiply it with grams Cu (because that is also what the question asks for) over 1 mole Cu. After canceling and multiplying out what is left, you come to 127g Cu.)

2.  $27\text{g Cu} \times 1 \text{ mole Cu}/63.5\text{g Cu} \times 2 \text{ mole Cr}/3 \text{ mole Cu} \times 52\text{g Cr}/1 \text{ mole Cr} = 14.74\text{g Cr}$

("excess of CuSO<sub>4</sub>" is irrelevant to the problem you are solving, so you can cross that out. You are left only to work with 27g Cu, and need to find grams Cr. So you divide 27g Cu by its molar mass [from periodic table] and then multiply it by the mole to mole ratio of 2 moles Cr over 3 moles Cu. You put 3 moles Cu on the bottom to cancel it out with 1 mole Cu. Lastly, to cancel out 2 moles Cr, multiply it with the molar mass [52g] over 1 mole Cr and multiply out what is uncrossed.)

## Section 12 3 Limiting Reagent And Percent

**Kenneth W. Raymond**



## **Section 12 3 Limiting Reagent And Percent:**

## **Section 12 3 Limiting Reagent And Percent** Book Review: Unveiling the Magic of Language

In an electronic digital era where connections and knowledge reign supreme, the enchanting power of language has become more apparent than ever. Its ability to stir emotions, provoke thought, and instigate transformation is truly remarkable. This extraordinary book, aptly titled "**Section 12 3 Limiting Reagent And Percent**," published by a very acclaimed author, immerses readers in a captivating exploration of the significance of language and its profound effect on our existence. Throughout this critique, we shall delve to the book's central themes, evaluate its unique writing style, and assess its overall influence on its readership.

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