

# Limiting Reactant, Theoretical Yield, Actual Yield, & Percent Yield

$22.99 + 35.45 = 58.44$

24



- If we have 53.2 g of Na and 65.8 g of  $\text{Cl}_2$ , what are the limiting reactant and theoretical yield?

$$53.2\text{ g Na} \times \frac{1\text{ mol Na}}{22.99\text{ g Na}} \times \frac{2\text{ mol NaCl}}{2\text{ mol Na}} \times \frac{58.44\text{ g NaCl}}{1\text{ mol NaCl}} = 135.23$$

$$65.8\text{ g Cl}_2 \times \frac{1\text{ mol Cl}_2}{70.9\text{ g Cl}_2} \times \frac{2\text{ mol NaCl}}{1\text{ mol Cl}_2} \times \frac{58.44\text{ g NaCl}}{1\text{ mol NaCl}} = 110.5$$

## Section 12 3 Limiting Reagent And Percent Yoeld

**ML Morrison**



## **Section 12 3 Limiting Reagent And Percent Yoeld:**

## Reviewing **Section 12 3 Limiting Reagent And Percent Yoeld**: Unlocking the Spellbinding Force of Linguistics

In a fast-paced world fueled by information and interconnectivity, the spellbinding force of linguistics has acquired newfound prominence. Its capacity to evoke emotions, stimulate contemplation, and stimulate metamorphosis is actually astonishing. Within the pages of "**Section 12 3 Limiting Reagent And Percent Yoeld**," an enthralling opus penned by a highly acclaimed wordsmith, readers attempt an immersive expedition to unravel the intricate significance of language and its indelible imprint on our lives. Throughout this assessment, we shall delve into the book's central motifs, appraise its distinctive narrative style, and gauge its overarching influence on the minds of its readers.

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