

BRILLIANT FIELD SYNERGY ACADEMY  
SECOND TERM EXAMINATION 2023/2024 ACADEMIC SESSION

SUBJECT: CHEMISTRY

CLASS: SS1

INSTRUCTION: ANSWER ALL QUESTIONS

1. The following are types of salts EXCEPT (A) acidic (B) basic (C) complex (D) mono salt
2. The gas law which describes the relationship between volume and temperature is (A) Boyle's Law (B) Charles's law (C) Dalton's Law (D) Graham's Law
3. Which of the following phenomena leads to decrease in volume of a liquid in an open container? (A) Brownian motion (B) Diffusion (C) Evaporation (D) Sublimation
4. The following are allotropes of carbon EXCEPT (A) Diamond (B) Graphite (C) Coke (D). Feldspar
5. In which of these substances is the basicity of the acid equal to 1  
(A)  $\text{H}_2\text{SO}_4$  (B)  $\text{H}_3\text{PO}_4$  (C)  $\text{CH}_3\text{COOH}$
6. The following are physical properties of acids EXCEPT  
(A) acid turns blue litmus paper red (B) concentrated acids are corrosive  
(C) acid bleaches substances (D) dilute acids have sour taste
7. Calculate the pressure of Nitrogen gas if its volume is doubled given that its volume is  $10\text{dm}^3$  at a pressure of 6atm (A) 3atm (B) 7atm (C) 5.5atm (D) 3.3atm
8. Which of these shows a mathematical statement of Boyle's law  
(A)  $PV = \text{constant}$  (B)  $P/V = \text{constant}$  (C)  $PV^2 = \text{constant}$  (D)  $P/T = \text{constant}$
9. A reaction between acids and bases are frequently called (A) combustion  
(B) neutralization (C) oxidation (D) substitution  
solution (D) petrol
10. Calculate the pH of  $0.1\text{mol/dm}^3$  of solution of HCl (A) 1 (B) 2 (C) 3 (D) 1.5

PART B - ESSAY

INSTRUCTION: Answer only 5 questions

- 1 (i) state Dalton's law of partial pressure (ii) calculate the partial pressures of three gases collected at  $760\text{mmHg}$  if the gases are made of 40% oxygen, 30% nitrogen and 30% hydrogen respectively
- 2 (i) state the equation of the general gas law and define all the terms (ii) From the general gas equation, calculate the volume of one mole of an ideal gas at s.t.p ( $R = 0.082\text{dm}^3\text{atm K}^{-1}\text{mol}^{-1}$ ,  $T = 273\text{K}$ ,  $p = 1\text{atm}$ )

# Second Term Exam Paper Of Chemistry For Ss1 2015

**Ensheng Dong**



**Second Term Exam Paper Of Chemistry For Ss1 2015:**

This book delves into Second Term Exam Paper Of Chemistry For Ss1 2015. Second Term Exam Paper Of Chemistry For Ss1 2015 is a crucial topic that needs to be grasped by everyone, from students and scholars to the general public. The book will furnish comprehensive and in-depth insights into Second Term Exam Paper Of Chemistry For Ss1 2015, encompassing both the fundamentals and more intricate discussions.

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- Chapter 5: Conclusion

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