

7. Chromite is the main ore of chromium. The chromium-containing compound in chromite is FeCr_2O_4 . The percentage of chromium in a sample of chromite can be determined using the method below.

Step 1

A 5.28 g sample of chromite ore is heated with sodium peroxide, Na_2O_2 .



Water is added to the resulting mixture.

Na_2O_2 dissolves in the water forming a solution containing CO_3^{2-} ions.

Step 2

The mixture from Step 1 is filtered and the filtrate is made up to 1.00 dm³ in a volumetric flask.

A 25.0 cm³ sample of this alkaline solution is pipetted into a conical flask and an excess of aqueous potassium iodide is added.

- A redox reaction takes place between I^- ions, CO_3^{2-} ions and H_2O .
- In this reaction 1 mol CO_3^{2-} forms 1.5 mol I_2 .

Step 3

The resulting mixture is titrated with 0.100 mol dm⁻³ sodium thiosulfate, $\text{Na}_2\text{S}_2\text{O}_3$, to estimate the I_2 present.



The average titre of $\text{Na}_2\text{S}_2\text{O}_3$ was 26.5 cm³.

- (a) In Step 1, Na_2O_2 and NaFeO_2 react with water forming an alkaline solution containing a brown precipitate. This is not a redox reaction.

Write equations for:

- the reaction of Na_2O_2 with water
- the reaction of NaFeO_2 with water.



- (b) Calculate the percentage, by mass, of chromium in the ore.

Give your answer to one decimal place.

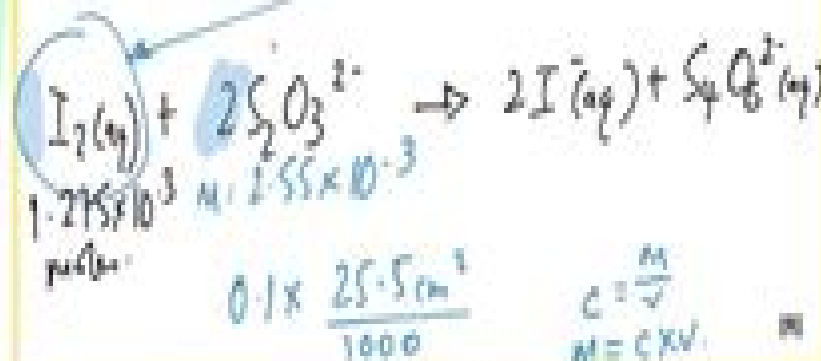
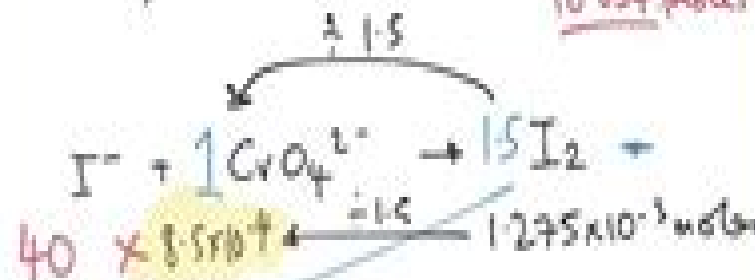
$$\% \text{ Cr} = \frac{\text{mass Cr}}{\text{mass ore}} \times 100$$

$$\text{mass} = \text{moles} \times \text{M.R.}$$

$$1.76 \quad 0.034 \times 52.0$$



0.034 moles



- (c) This part refers to Step 2 of the method.

is the redox reaction between I^- ions, CO_3^{2-} ions and H_2O .

- CO_3^{2-} ions are reduced to carbonate ions, CO_3^{2-} .
- I^- ions are oxidised to make I_2 .
- Construct an overall equation for the redox reaction and write half equations for the oxidation and reduction.

Overall equation:

Half equations:

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1. Understanding the eBook Ocr Chemistry F325 Stretch And Challenge
 - The Rise of Digital Reading Ocr Chemistry F325 Stretch And Challenge
 - Advantages of eBooks Over Traditional Books
2. Identifying Ocr Chemistry F325 Stretch And Challenge
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Ocr Chemistry F325 Stretch And Challenge
 - User-Friendly Interface
4. Exploring eBook Recommendations from Ocr Chemistry F325 Stretch And Challenge

- Personalized Recommendations
- Ocr Chemistry F325 Stretch And Challenge User Reviews and Ratings
- Ocr Chemistry F325 Stretch And Challenge and Bestseller Lists
- 5. Accessing Ocr Chemistry F325 Stretch And Challenge Free and Paid eBooks
 - Ocr Chemistry F325 Stretch And Challenge Public Domain eBooks
 - Ocr Chemistry F325 Stretch And Challenge eBook Subscription Services
 - Ocr Chemistry F325 Stretch And Challenge Budget-Friendly Options
- 6. Navigating Ocr Chemistry F325 Stretch And Challenge eBook Formats
 - ePub, PDF, MOBI, and More
 - Ocr Chemistry F325 Stretch And Challenge Compatibility with Devices
 - Ocr Chemistry F325 Stretch And Challenge Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Ocr Chemistry F325 Stretch And Challenge
 - Highlighting and Note-Taking Ocr Chemistry F325 Stretch And Challenge
 - Interactive Elements Ocr Chemistry F325 Stretch And Challenge
- 8. Staying Engaged with Ocr Chemistry F325 Stretch And Challenge
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Ocr Chemistry F325 Stretch And Challenge
- 9. Balancing eBooks and Physical Books Ocr Chemistry F325 Stretch And Challenge
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Ocr Chemistry F325 Stretch And Challenge
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Ocr Chemistry F325 Stretch And Challenge
 - Setting Reading Goals Ocr Chemistry F325 Stretch And Challenge
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Ocr Chemistry F325 Stretch And Challenge

- Fact-Checking eBook Content of Ocr Chemistry F325 Stretch And Challenge
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
- Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
- Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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