Question	Answer	Marks	Guidance
number			A STATE OF THE STA
6 (c)	$M(NaNO_3) = 85.0 \text{ g mol}^{-1}$	B1	
	n(NaNO ₃) = 0.04 mol	81	
	$n(O_2) = 0.02 \text{ mol}$	B1	
	Volume of O ₂ = 0.0200 × 24 000 = 480 cm ³	B1	
7 (a)	0.0250 × 23.0 = 0.575 g	81	
7 (b)	2Na(s) + 2H ₂ O(l) → 2NaOH(aq) + H ₂ (g)	81	
7 (c)	n(H ₂) = 0.0125 mol	B1	
	Volume H ₂ = 0.0125 × 24 000 = 300 cm ³	B1	
7 (d) (i)	$n(NaOH) = n(Na) = 0.0250 = c \times 50/1000$ $c = 0.500 \text{ mol dm}^{-3}$	B1	
7 (d) (ii)	$c = 0.500 \times 40.0 = 20.0 \mathrm{g}\mathrm{dm}^{-3}$	B1	
8 (a)	58.5 g	B1	
8 (b)	$n(Cl_2) = (2.5 \times 10^5)/24 = 1.04 \times 10^5 \text{ mol}$	B1	
	$n(NaOH) = 2 \times 1.04 \times 10^{8} = 2.08 \times 10^{8} \text{ mol}$	B1	
	$2.08 \times 10^8 = 4.00 \times V \text{ (in dm}^3\text{)}$ $\therefore V = 5.2 \times 10^7 \text{ dm}^3$	B1	
9 (a) (i)	C: H = 54.55/12.0: 9.09/1.0: 36.36/16.0 = 4.55 : 9.09: 2.27	B1	
	Empirical formula ≈ C ₂ H ₄ O	B1	
9 (a) (ii)	pV = nRT	B1	
	$n = \frac{(103 \times 10^3) \times (72.0 \times 10^{-6})}{8.314 \times 373} =$	B1	
	0.002 39 mol	B1	
	$M = \frac{0.2103}{0.00239} = 88.0$	B1	
9 (a) (iii)	Molecular formula = $C_2H_4O \times 88/44 = C_4H_8O_2$	B1	
10 (a) (i)	Al ³¹	81	
	SO ₄ 2-	B1	

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- 1. Understanding the eBook Ocr Chemistry As Answers
 - The Rise of Digital Reading Ocr Chemistry As Answers
 - Advantages of eBooks Over Traditional Books
- 2. Identifying Ocr Chemistry As Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
- 3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Ocr Chemistry As Answers
 - User-Friendly Interface
- 4. Exploring eBook Recommendations from Ocr Chemistry As Answers
 - Personalized Recommendations
 - Ocr Chemistry As Answers User Reviews and Ratings
 - Ocr Chemistry As Answers and Bestseller Lists

- 5. Accessing Ocr Chemistry As Answers Free and Paid eBooks
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 - Highlighting and Note-Taking Ocr Chemistry As Answers
 - Interactive Elements Ocr Chemistry As Answers
- 8. Staying Engaged with Ocr Chemistry As Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Ocr Chemistry As Answers
- 9. Balancing eBooks and Physical Books Ocr Chemistry As Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Ocr Chemistry As Answers
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Ocr Chemistry As Answers
 - Setting Reading Goals Ocr Chemistry As Answers
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Ocr Chemistry As Answers
 - Fact-Checking eBook Content of Ocr Chemistry As Answers
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
- Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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