# Economics 1 OCR

Oxford Cambridge and RSA

# Monday 23 May 2022 – Afternoon

#### A Level Economics

H460/01 Microeconomics

**JUNE 2022** 

Time allowed: 2 hours

a scientific or graphical calculator	
Please write clearly in black ink	. Do not write in the barcodes.  Candidate number
First name(s) Last name	

#### INSTRUCTIONS

- Use black ink. You can use an HB pencil, but only for graphs and diagrams.
- Write your answer to each question in the space provided. If you need extra space use the lined pages at the end of this booklet. The question numbers must be clearly shown.
- Answer all the questions in Section A, one question in Section B and one question in Section C.

#### INFORMATION

- The total mark for this paper is 80.
- The marks for each question are shown in brackets [ ].
- Quality of extended response will be assessed in questions marked with an asterisk (\*).
- This document has 20 pages.

#### ADVICE

Read each question carefully before you start your answer.

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web the heat of reaction  $\Delta$ hsoln is written after the products in units of kj mol rxn for example sodium hydroxide dissolves exothermically releasing 44 2 kilojoules per mole dissolved this process is represented as naoh s na aq oh aq  $\Delta$ hsoln 44 2 kj mol rxn write equations to similarly represent the dissolving process for

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web hand warmers are usually placed inside the gloves to keep your hands warm the heat keeping your hands warm is formed as a result of an exothermic reaction which occurs in a small packet hand warmers have 5 main ingredients in them salt carbon polypropylene vermiculite water and iron

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water this chemical hand warmer is reusable

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web in this lab you are challenged to use chemistry to determine which ionic compound makes the best hand warmer in this experiment you will collect data that will allow you to calculate the heat of solution  $\Delta$ hsoln or the amount of energy released or absorbed in kj mol of the substance being dissolved

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web mar 4 2014 contains water the hand warmer is struck in a manner that ruptures the inner pouch releasing the ionic salt into the water of the outer pouch the salt dissolves and the water warms this type of hand warmer tends to produce a more vigorous heat than the dry powder type of hand warmer but does not produce heat for quite as long

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