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## Assessment

## Quiz

## Section: Calculating Quantities in Reactions

In the space provided, write the letter of the term or phrase that best answers the question.

- \_\_\_\_\_ 1. The chemical equation  $\text{P}_4\text{O}_{10} + 6\text{H}_2\text{O} \rightarrow 4\text{H}_3\text{PO}_4$  can be correctly interpreted as
- 1 molecule of  $\text{P}_4\text{O}_{10} + 6$  molecules of  $\text{H}_2\text{O}$  form 4 molecules of  $\text{H}_3\text{PO}_4$ .
  - 1 mol  $\text{P}_4\text{O}_{10} + 6$  mol  $\text{H}_2\text{O}$  reacts to form 4 mol  $\text{H}_3\text{PO}_4$ .
  - 1 g  $\text{P}_4\text{O}_{10} + 6$  g  $\text{H}_2\text{O}$  produces 4 g  $\text{H}_3\text{PO}_4$ .
  - Both (a) and (b)
- \_\_\_\_\_ 2. Stoichiometry is the branch of chemistry that deals with the \_\_\_\_\_ between elements and compounds as reactants and products in a chemical reaction.
- bonding
  - energy transfers
  - numerical relationship
  - physical interactions
- \_\_\_\_\_ 3. How many mole ratios can be correctly obtained from the following chemical equation?
- $$\text{Na}_2\text{S} + \text{Cd}(\text{NO}_3)_2 \rightarrow 2\text{NaNO}_3 + \text{CdS}$$
- 1
  - 4
  - 12
  - 16
- \_\_\_\_\_ 4. Molar mass can be used to relate
- mass of reactant to moles of reactant.
  - moles of product to mass of product.
  - mass of product to moles of product.
  - All of the above
- \_\_\_\_\_ 5. In the chemical reaction described by the equation
- $$4\text{Fe}(s) + 3\text{O}_2(g) \rightarrow 2\text{Fe}_2\text{O}_3(s),$$
- the mole ratio of iron(III) oxide to iron is
- 1:1.
  - 1:2.
  - 2:3.
  - 4:1.

# Quiz Section Calculating Quantities In Reactions

**DP Hallahan**



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