## **Balancing Act**

Name

Atoms are not <u>created</u> or <u>destroyed</u> during a chemical reaction.

Scientists know that there must be the <u>same</u> number of atoms on each <u>side</u> of the <u>reaction</u>. To balance the chemical equation, you must add <u>coefficients</u> in front of the chemical formulas in the equation. You cannot <u>add</u> or <u>remove</u> subscripts.

- 1) Determine number of atoms for each element
- 2) Pick an element that is not equal on both sides of the equation.
- 3) Add coefficients in front of the element and adjust your counts.
- Continue adding coefficients until you get the same number of elements on each side.

#### Example:

$$2$$
 Mg +  $O_2$   $\rightarrow$   $2$  MgO  
Mg =  $2$  Mg =  $2$   
O =  $2$  O =  $2$ 

#### Try these:

$$N_2$$
 + 3  $H_2$   $\rightarrow$  2  $NH_3$   
 $N = 2$   
 $H = 6$   $H = 6$ 

$$2 H_2O_2 \rightarrow 2 H_2O + O_2$$
 $H = 4 H = 4$ 
 $O = 4 O = 4$ 

# **Sciencespot Net Balancing Act Practice Answer Key**

**JA Banks** 

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