

Section 1 Review page 118

1. **Describe Democritus's atomic theory.**
Everything in the universe is made up of atoms.
2. **Summarize the main ideas of Dalton's theory.**
 - a. Elements are made up of tiny unique particles called atoms.
 - b. Atoms cannot be divided.
 - c. Atoms of the same element are identical.
 - d. Atoms of different elements can join to form molecules.
3. **Explain why Dalton's theory was more successful than Democritus's theory.**
Dalton had experimental evidence to support his theory, while Democritus did not.
4. **Compare Thomson's atomic model with Rutherford's atomic model.**
Thomson's model- electrons were embedded in a sphere of positive charge
Rutherford's model- positive charge is concentrated in the nucleus and electrons surrounded the nucleus.
5. **Analyzing elements**
 - a. **How did the cathode-ray tube experiment lead to the conclusion that atoms contain electrons?** The air was removed from the tube and the beam of negative particles was observed. This suggests that atoms are made up of smaller units with negative charges.
 - b. **How did the gold-foil experiment lead to the conclusion that the atom has a nucleus?** the positive particles were deflected at large angles and therefore this suggests that a positive charge is concentrated in the center of the atom. (nucleus)
6. **Making Inferences: Does the term *indivisible* still describe the atom? Explain.**
No, the atom is not indivisible. It is made up of smaller parts such as the electron.

Section 2 review page 127

1. **List the charge, mass, and location of each of the three subatomic particles found in atoms.** The proton is positive, the neutron is neutral and electrons are negative. Protons and neutrons are found in the nucleus of an atom. The mass of a proton 1850 times greater than that of an electron.
2. **Explain how you can use an atom's mass number and atomic number to determine the number of protons, electrons, and neutrons to determine the number of protons, electrons, and neutrons in the atom.** The atomic number equals the number of protons. The number of electrons equals the number of protons. To find the number of neutrons subtract the atomic number from the mass number.
3. **Identify the subatomic particle used to define an element, and explain why this particle is used.** The proton defines the element's number because the each element always has the same number of protons, while the number of neutrons and electrons can vary.
4. **Explain why the masses of atoms of the same element may differ.** Because they can have different number of neutrons. For example Hydrogen (H) H, H₁Deuterium, H₂ Tritium. These are called isotopes.

Science Spectrum Physical Science Chapter 4 Section Reviews Answers

Kathleen Armour



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Science and Technology, Committee on Review of the U.S. Department of Defense, Air and Space Systems Science and Technology Program, 2001-09-05 Since the mid 1940s when Vannevar Bush and Theodore von Karman wrote Science the Endless Frontier and Toward New Horizons respectively there has been a consensus that strong Department of Defense support of science and technology S T is important to the security of the United States During the Cold War as it faced technologically capable adversaries whose forces potentially outnumbered U S forces the United States relied on a strong defense S T program to support the development of technologically superior weapons and systems that would enable it to prevail in the event of conflict Since the end of the Cold War the United States has relied on its technological superiority to maintain a military advantage while at the same time reducing the size of its forces Over the past half century creating and maintaining a technologically superior military capability have become fundamental to U S national security strategy and investment in S T has become a basic component of the defense budget In late 1998 Congress asked the Secretary of Defense to conduct a study in cooperation with the National Research Council NRC on the S T base of the U S Department of Defense DoD Congress was particularly concerned about areas of the S T program related to air systems space systems and supporting information systems Its concern was based on the Air Force s reduction of its S T program from the largest of the three military service programs to the smallest Congress also wanted to ensure that the Air Force maintained an appropriately sized S T workforce In late 1999 the Deputy Under Secretary of Defense for Science and Technology asked the NRC to conduct a study to explore these issues

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