

Types of Chemical Reactions Worksheet #1

11-Chemistry

For each of the following equations, identify what kind of reaction it represents: double replacement, single replacement, decomposition, or synthesis (composition).

- 1) $2\text{Mg} + \text{O}_2 \rightarrow 2\text{MgO}$ **Synthesis**
- 2) $\text{CaI}_2 + \text{Cl}_2 \rightarrow \text{CaCl}_2 + \text{I}_2$ **Single Replacement**
- 3) $\text{HNO}_3 + \text{AlCl}_3 \rightarrow \text{Al}(\text{OH})_3 + 3\text{HCl}$ **Double Replacement**
- 4) $\text{C} + \text{O}_2 \rightarrow \text{CO}_2$ **Synthesis**
- *5) $\text{Ca}(\text{OH})_2 + 2\text{HCl} \rightarrow \text{CaCl}_2 + 2\text{H}_2\text{O}$ **Double Replacement**
- 6) $2\text{KClO}_3 \rightarrow 2\text{KCl} + 3\text{O}_2$ **Decomposition**
- 7) $3\text{CuSO}_4 + 2\text{Al} \rightarrow \text{Al}_2(\text{SO}_4)_3 + 3\text{Cu}$ **Single Replacement**
- 8) $\text{Na}_2\text{S} + 2\text{AgNO}_3 \rightarrow 2\text{NaNO}_3 + \text{Ag}_2\text{S}$ **Double Replacement**
- 9) $3\text{H}_2 + \text{N}_2 \rightarrow 2\text{NH}_3$ **Synthesis**

*10) $2\text{Na} + 2\text{H}_2\text{O} \rightarrow 2\text{NaOH} + \text{H}_2$ **Single Replacement**

*Remember: "H₂O" can also be thought of as "HOH"

Complete each of the following equations as needed to make it the type of reaction indicated. Be sure to write each formula correctly.

- 11) Double replacement: $\text{Na}_2\text{CrO}_4 + \text{PbCl}_2 \rightarrow 2\text{NaCl} + \text{PbCrO}_4$
- 12) Single replacement: $\text{Cl}_2 + 2\text{NaBr} \rightarrow 2\text{NaCl} + \text{Br}_2$
- 13) Decomposition: $\text{Mg}(\text{ClO}_3)_2 \rightarrow \text{MgCl}_2 + 3\text{O}_2$
- 14) Synthesis: $2\text{H}_2 + \text{O}_2 \rightarrow 2\text{H}_2\text{O}$
- 15) Double replacement: $3\text{Ca}(\text{OH})_2 + 2\text{FeCl}_3 \rightarrow 2\text{Fe}(\text{OH})_3 + 3\text{CaCl}_2$
- 16) Single replacement: $\text{Fe} + \text{Cu}(\text{NO}_3)_2 \rightarrow \text{Fe}(\text{NO}_3)_2 + \text{Cu}$ [Assume Fe²⁺]
- 17) Decomposition: $2\text{Hg}_2\text{O} \rightarrow 4\text{Hg} + \text{O}_2$
- 18) Synthesis: $\text{S} + \text{O}_2 \rightarrow \text{SO}_2$
- 19) Double replacement: $\text{AgNO}_3 + \text{KI} \rightarrow \text{AgI} + \text{KNO}_3$
- 20) Single replacement: $\text{Cu} + 2\text{AgNO}_3 \rightarrow 2\text{Ag} + \text{Cu}(\text{NO}_3)_2$ [Copper (II) is used here]

Science Skills Chemical Reactions Answers

Amar Bhutani



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Harsch, Siegbert Schmid, 2011-10-28 For everybody teaching chemistry or becoming a chemistry teacher the authors provide a practice oriented overview with numerous examples from current chemical education including experiments models and exercises as well as relevant results from research on learning and teaching With their proven concept the authors cover classical topics of chemical education as well as modern topics such as every day life chemistry students' misconceptions the use of media or the challenges of motivation This is the completely revised and updated English edition of a highly successful German title **Chemistry: The Central Science** Theodore L. Brown, H. Eugene LeMay Jr., Bruce E. Bursten, Catherine

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Beginning Science Teaching Chris Dawson, 1991

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Contents list 1 Where are we now and where are we going 2 How do we map matter 3 Who are we 4 How can we find out 5 How does our planet work 6 How do we respond to our world

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Table of Contents Science Skills Chemical Reactions Answers

1. Understanding the eBook Science Skills Chemical Reactions Answers
 - The Rise of Digital Reading Science Skills Chemical Reactions Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Science Skills Chemical Reactions Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Science Skills Chemical Reactions Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Science Skills Chemical Reactions Answers

- Personalized Recommendations
 - Science Skills Chemical Reactions Answers User Reviews and Ratings
 - Science Skills Chemical Reactions Answers and Bestseller Lists
5. Accessing Science Skills Chemical Reactions Answers Free and Paid eBooks
 - Science Skills Chemical Reactions Answers Public Domain eBooks
 - Science Skills Chemical Reactions Answers eBook Subscription Services
 - Science Skills Chemical Reactions Answers Budget-Friendly Options
 6. Navigating Science Skills Chemical Reactions Answers eBook Formats
 - ePub, PDF, MOBI, and More
 - Science Skills Chemical Reactions Answers Compatibility with Devices
 - Science Skills Chemical Reactions Answers Enhanced eBook Features
 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Science Skills Chemical Reactions Answers
 - Highlighting and Note-Taking Science Skills Chemical Reactions Answers
 - Interactive Elements Science Skills Chemical Reactions Answers
 8. Staying Engaged with Science Skills Chemical Reactions Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Science Skills Chemical Reactions Answers
 9. Balancing eBooks and Physical Books Science Skills Chemical Reactions Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Science Skills Chemical Reactions Answers
 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
 11. Cultivating a Reading Routine Science Skills Chemical Reactions Answers
 - Setting Reading Goals Science Skills Chemical Reactions Answers
 - Carving Out Dedicated Reading Time
 12. Sourcing Reliable Information of Science Skills Chemical Reactions Answers

- Fact-Checking eBook Content of Science Skills Chemical Reactions Answers
- Distinguishing Credible Sources

13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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