## Study Guide: Thermal Energy

#### I. What is Thermal Energy?

- Definition: The total kinetic and potential energy of all the particles in a substance.
- Relationship to Particle Motion: The faster the particles move, the more thermal energy they have.
- Factors Affecting Thermal Energy:
  - Temperature
  - 6-September
  - State of matter (solid, liquid, gas)

#### III. Temperature

- Definition: A measure of the average kinetic energy of the particles in a substance.
- Units: Degrees Celsius (\*C), degrees Fahrenheit (\*F), Kelvin (kr)
- Relationship to Thermal Energy: Higher temperature means higher average kinetic energy, which generally means more thermal energy.

#### Mt. Heat

- Definition: The transfer of thermal energy from a warmer object to a cooler object.
- Direction of Heat Flows Heat always flows from hot to cold until thermal equilibrium is reached.
- Unit s: (oules (i), calories (cal).
  - Methods of Heat Transfer:
    - Conductions Heat transfer through direct contact.
      - · Example: Touching a hot stove.
    - Convection: Heat transfer through the movement of fluids (liquids or gases).
      - Example: Boiling water in a pot.
    - Radiation: Heat transfer through electromagnetic
      - Example: Feeling the heat from the sun.

#### IV. Thermal Properties of Matter

- Thermal Expansion and Contraction: Most substances expand when heated and contract when cooled.
  - Explanation: Particles move farther apart when heated and closer together when cooled.

# **Science Thermal Energy Study Guide**

Barron's Educational Series, Brian Stewart

## **Science Thermal Energy Study Guide:**

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