

# CHEMISTRY



**Periodic Table of Elements**

**Properties of the Periodic Table**

- The periodic table is a tabular arrangement of the chemical elements, organized on the basis of their atomic number, electron configuration, and recurring chemical properties.
- The elements are arranged in rows (periods) and columns (groups).
- The periodic table is a fundamental tool in chemistry, used to predict the properties of elements and to understand the relationships between them.

**Atomic Trends**

- Atomic size increases from top to bottom and from right to left.
- Ionization energy increases from bottom to top and from left to right.
- Electronegativity increases from bottom to top and from left to right.

**Atomic Structure**

Atoms are made up of three subatomic particles: electrons, protons, and neutrons. Electrons are negatively charged, protons are positively charged, and neutrons are neutral.

The atomic number of an element is the number of protons in its nucleus. The mass number is the sum of the number of protons and neutrons in the nucleus.

**Atomic Structure: Protons & Electrons**

Protons and electrons are present in all atoms. The number of protons is equal to the number of electrons, making the atom electrically neutral.

The mass of a proton is approximately 1836 times that of an electron.

**Properties of Protons and Electrons**

Property	Proton	Electron
Charge	Positive (+1)	Negative (-1)
Mass	1.6726 × 10 <sup>-27</sup> kg	9.1091 × 10 <sup>-31</sup> kg
Location	Nucleus	Outside the nucleus

**Atomic Structure: Neutrons**

Neutrons are present in the nucleus of an atom. They are neutral particles with a mass slightly greater than that of a proton.

The number of neutrons in an atom is given by the mass number minus the atomic number.

**Atomic Structure: Nucleus**

The nucleus is the central part of an atom, containing protons and neutrons. It is extremely small and dense.

The size of the nucleus is on the order of 10<sup>-14</sup> m.

**Atomic Structure: Electrons**

Electrons are present in the outer regions of an atom, called the electron cloud. They move in orbits around the nucleus.

The number of electrons in an atom is equal to the number of protons.

**Atomic Structure: Atomic Number**

The atomic number (Z) is the number of protons in the nucleus of an atom. It is a unique identifier for each element.

The atomic number determines the chemical properties of an element.

**Atomic Structure: Mass Number**

The mass number (A) is the sum of the number of protons and neutrons in the nucleus of an atom.

The mass number is a whole number, while the atomic number is not necessarily a whole number.

**Atomic Structure: Isotopes**

Isotopes are atoms of the same element that have the same atomic number but different mass numbers.

Isotopes are used in various applications, including medicine and industry.

**Atomic Structure: Atomic Mass**

The atomic mass is the mass of an atom, expressed in atomic mass units (amu).

The atomic mass is a weighted average of the masses of the isotopes of an element.

**Atomic Structure: Atomic Size**

The atomic size is the radius of an atom, measured in meters (m).

The atomic size increases from top to bottom and from right to left in the periodic table.

**Atomic Structure: Atomic Density**

The atomic density is the mass of an atom divided by its volume.

The atomic density is a measure of how tightly packed the particles in an atom are.

**Atomic Structure: Atomic Stability**

The atomic stability is the ability of an atom to resist changes in its structure.

Atoms with a full outer shell of electrons are generally more stable than those with a partially filled shell.

**Atomic Structure: Atomic Spectra**

The atomic spectrum is the pattern of lines of different wavelengths that are emitted or absorbed by an atom.

The atomic spectrum is a unique identifier for each element.

**Atomic Structure: Atomic Applications**

Atomic structure is the foundation of chemistry and physics. It helps us understand the behavior of matter and the forces that hold it together.

Atomic structure is also used in many practical applications, such as in the development of new materials and the treatment of diseases.

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