

Quick Lab

Purpose To use a graph to identify period and group trends

Materials

- graph paper
- pencil



Periodic Trends in Ionic Radii

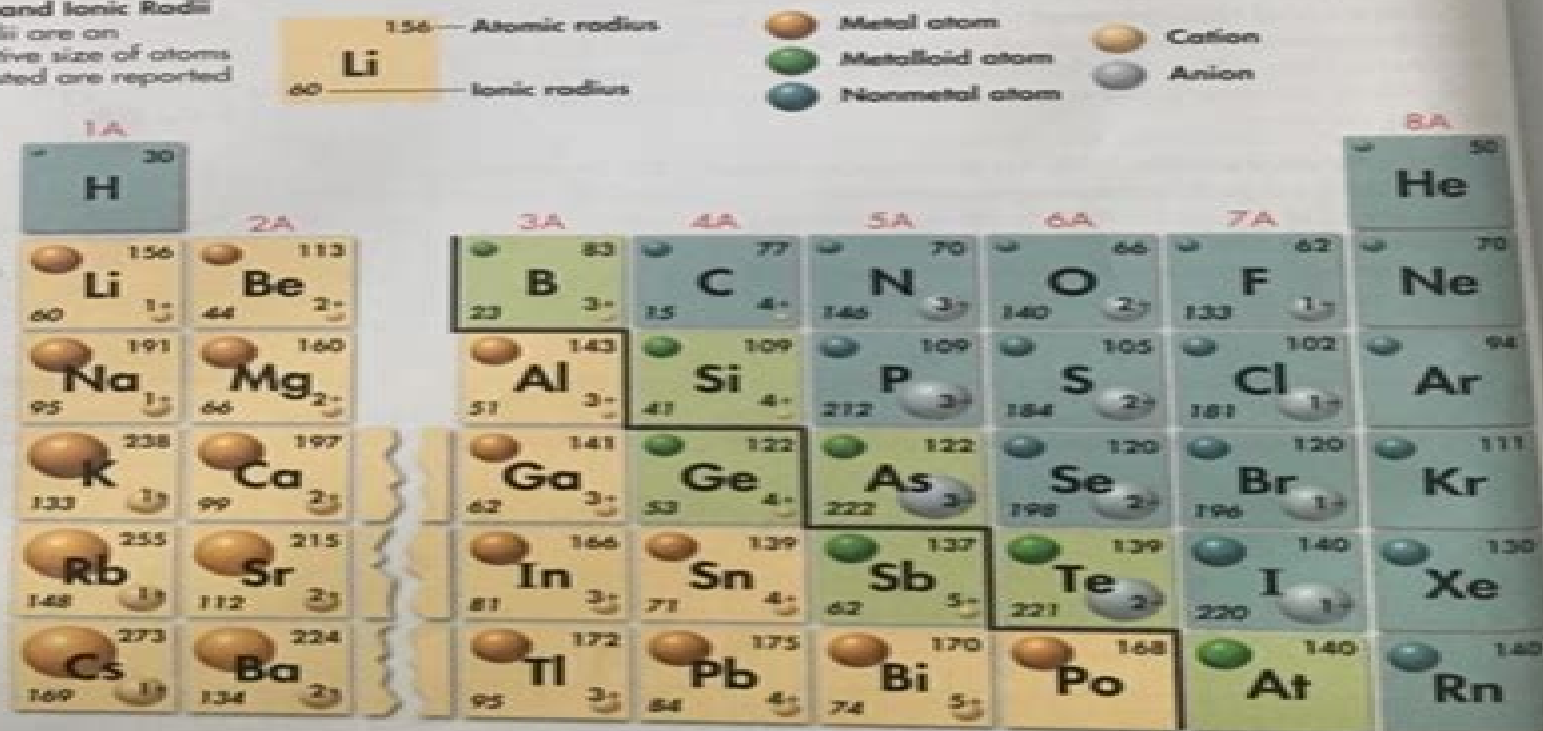
Procedure

Use the data presented in Figure 6.23 to plot ionic radius versus atomic number.

Analyze and Conclude

- 1. Compare** How does the size change when an atom forms a cation and when an atom forms an anion?
- 2. Describe** How do the ionic radii vary within a group of metals? How do they vary within a group of nonmetals?
- 3. Describe** What is the shape of a portion of the graph that corresponds to one period?
- 4. Compare and Contrast** Is the trend across a period similar or different for Periods 2, 3, 4, and 5?
- 5. Explain** Propose explanations for the trends you have described for ionic radii within groups and across periods.

Figure 6.23 Atomic and Ionic Radii
Atomic and ionic radii are an indication of the relative size of atoms and ions. The data listed are reported in picometers (pm).



Quick Lab Periodic Table In Ionic Radii

**Norbert Adolph Lange, Gordon M.
Forker**



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Table of Contents Quick Lab Periodic Table In Ionic Radii

1. Understanding the eBook Quick Lab Periodic Table In Ionic Radii
 - The Rise of Digital Reading Quick Lab Periodic Table In Ionic Radii
 - Advantages of eBooks Over Traditional Books
2. Identifying Quick Lab Periodic Table In Ionic Radii
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Quick Lab Periodic Table In Ionic Radii
 - User-Friendly Interface
4. Exploring eBook Recommendations from Quick Lab Periodic Table In Ionic Radii
 - Personalized Recommendations
 - Quick Lab Periodic Table In Ionic Radii User Reviews and Ratings
 - Quick Lab Periodic Table In Ionic Radii and Bestseller Lists
5. Accessing Quick Lab Periodic Table In Ionic Radii Free and Paid eBooks
 - Quick Lab Periodic Table In Ionic Radii Public Domain eBooks
 - Quick Lab Periodic Table In Ionic Radii eBook Subscription Services
 - Quick Lab Periodic Table In Ionic Radii Budget-Friendly Options
6. Navigating Quick Lab Periodic Table In Ionic Radii eBook Formats
 - ePub, PDF, MOBI, and More
 - Quick Lab Periodic Table In Ionic Radii Compatibility with Devices
 - Quick Lab Periodic Table In Ionic Radii Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Quick Lab Periodic Table In Ionic Radii
 - Highlighting and Note-Taking Quick Lab Periodic Table In Ionic Radii
 - Interactive Elements Quick Lab Periodic Table In Ionic Radii
8. Staying Engaged with Quick Lab Periodic Table In Ionic Radii

- Joining Online Reading Communities
- Participating in Virtual Book Clubs
- Following Authors and Publishers Quick Lab Periodic Table In Ionic Radii
- 9. Balancing eBooks and Physical Books Quick Lab Periodic Table In Ionic Radii
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Quick Lab Periodic Table In Ionic Radii
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Quick Lab Periodic Table In Ionic Radii
 - Setting Reading Goals Quick Lab Periodic Table In Ionic Radii
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Quick Lab Periodic Table In Ionic Radii
 - Fact-Checking eBook Content of Quick Lab Periodic Table In Ionic Radii
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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