

½CHEM 101 Worksheet, Chapter 6: Thermochemistry

Multiple choice questions 1-10 (page 1): circle the correct answer(A-E)

Show your work questions 11-13 (pages 2-3): give the complete answer in the space provided

R = 8.314 J/(mol.K) = 0.08206 L.atm/(mol.K) 1 cal = 4.184 J 0 °C = 273 K

1. Calculate the heat effect q (in kJ) when 112 g of iron (Fe(s)) is heated from 20 °C to 300 °C. The specific heat of Fe(s) = 0.48 J/g.K

- A. +6.53 kJ B. +29.7 kJ C. -15.5 kJ D. -29.7 kJ E. +15.5 kJ

2. The dissolution reaction for ammonium chloride: $\text{NH}_4\text{Cl(s)} \rightarrow \text{NH}_4^+(\text{aq}) + \text{Cl}^-(\text{aq})$ is endothermic. When ammonium chloride is dissolved in water in an insulated (adiabatic) system, the temperature will:

- A. increase B. decrease C. stay the same

3. Consider the reaction: $\text{CaCO}_3(\text{s}) \rightarrow \text{CaO(s)} + \text{CO}_2(\text{g})$. $\Delta H_{\text{rxn}} = +178 \text{ kJ}$ What is the enthalpy change when 20.0 kg CaO is formed?

- A. $+6.36 \times 10^4 \text{ kJ}$ B. $-6.36 \times 10^4 \text{ kJ}$ C. $+6.36 \times 10^4 \text{ J}$ D. 63.6 kJ E. +178 kJ

4. Consider the reaction: $2\text{NO} + \text{O}_2 \rightarrow 2\text{NO}_2$. $\Delta H_{\text{rxn}} = -114.6 \text{ kJ}$ As always, the ΔH is "for the reaction as written"! Calculate the enthalpy change ΔH , when 12.6 g NO_2 is produced.

- A. -31.4 kJ B. +31.4 kJ C. -15.7 kJ D. +15.7 kJ E. none of these

5. How much heat is associated with the reaction that produces 256.1 g of ammonia, NH_3 ?

$\text{N}_2(\text{g}) + 3 \text{H}_2(\text{g}) \rightarrow 2 \text{NH}_3(\text{g})$ $\Delta H^\circ_{\text{rxn}} = -92.45 \text{ kJ}$

- A. 92.45 kJ of heat is released B. 695.1 kJ of heat is released C. 184.9 kJ of heat is absorbed
D. 184.9 kJ of heat is released E. 92 kJ of heat is released

6. Which two of the following are **not** formation reactions?

1. $\frac{1}{2}\text{N}_2(\text{g}) + \frac{1}{2}\text{H}_2(\text{g}) \rightarrow \text{NH}_3(\text{g})$ 2. $\text{C(s)} + \text{O}_2(\text{g}) \rightarrow \text{CO}_2(\text{g})$ 3. $2\text{SO}_2(\text{g}) + \text{O}_2(\text{g}) \rightarrow 2\text{SO}_3(\text{g})$
4. $2\text{Fe(s)} + \frac{1}{2}\text{O}_2(\text{g}) \rightarrow \text{Fe}_2\text{O}_3(\text{s})$ 5. $\frac{1}{2}\text{H}_2(\text{g}) + \frac{1}{2}\text{I}_2(\text{g}) \rightarrow \text{HI(g)}$

- A. 1 and 2 B. 3 and 4 C. 2 and 3 D. 4 and 5 E. 3 and 5

7. Consider the thermochemical reaction equation: $2\text{CO(g)} + \text{O}_2(\text{g}) \rightarrow 2\text{CO}_2(\text{g})$ $\Delta H = -198 \text{ kJ}$

What is the enthalpy change for the reaction: $\text{CO}_2(\text{g}) \rightarrow \text{CO(g)} + \frac{1}{2}\text{O}_2(\text{g})$

- A. +198 kJ B. -99 kJ C. +99 kJ D. +396 kJ E. -198 kJ

8. Calculate the enthalpy change for the reaction: $2\text{KClO}_3 \rightarrow 2 \text{KCl} + \text{O}_2(\text{g})$

Standard enthalpy of formation values: $\Delta H_f^\circ(\text{KClO}_3) = -389 \text{ kJ/mol}$ $\Delta H_f^\circ(\text{KCl}) = -436 \text{ kJ/mol}$

- A. -94 kJ B. -47 kJ C. +94 kJ D. +47 kJ
E. Cannot be calculated because $\Delta H_f^\circ(\text{O}_2(\text{g}))$ is not given

9. Given the following two thermochemical equations:

1. $2\text{H}_2(\text{g}) + \text{O}_2(\text{g}) \rightarrow 2\text{H}_2\text{O(l)}$ $\Delta H = -572 \text{ kJ}$

2. $\text{H}_2(\text{g}) + \text{O}_2(\text{g}) \rightarrow \text{H}_2\text{O}_2(\text{l})$ $\Delta H = -188 \text{ kJ}$

What is the enthalpy change for the reaction: $2\text{H}_2\text{O}_2(\text{l}) \rightarrow 2\text{H}_2\text{O(l)} + \text{O}_2(\text{g})$?

- A. -98 kJ B. +98 kJ C. +196 kJ D. -196 kJ E. -670 kJ

10. Which of the following statements is correct for a combustion reaction in an adiabatic "bomb" calorimeter?

- A. the pressure is constant B. the temperature is constant C. the volume is constant
D. the calorimeter constant's C units are J/g.K E. $\Delta T(\text{K}) = \Delta T(^{\circ}\text{C}) + 273.15$

Show your work for the following questions.

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Wolfgang Guggemos



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Internal Combustion Engines Shyam K. Agrawal, 2006 Salient Features The New Edition Is A Thoroughly Revised Version Of The Earlier Edition And Presents A Detailed Exposition Of The Basic Principles Of Design Operation And Characteristics Of Reciprocating I C Engines And Gas Turbines Chemistry Of Combustion Engine Cooling And Lubrication Requirements Liquid And Gaseous Fuels For Ic Engines Compressors Supercharging And Exhaust Emission Its Standards And Control Thoroughly Explained Jet And Rocket Propulsion Alternate Potential Engines Including Hybrid Electric And Fuel Cell Vehicles Are Discussed In Detail Chapter On Ignition System Includes Electronic Injection Systems For Si And Ci Engines 150 Worked Out Examples Illustrate The Basic Concepts And Self Explanatory Diagrams Are Provided Throughout The Text More Than 200 Multiple Choice Questions With Answers A Good Number Of Review Questions Numerical With Answers For Practice Will Help Users In Preparing For Different Competitive Examinations With These Features The Present Text Is Going To Be An Invaluable One For Undergraduate Mechanical Engineering Students And Amie Candidates **IIT Chemistry-II** , **Thermochemistry** Mr. Ajeet Singh, 2018-05-16 Introduces thermodynamic principles related to chemical processes including enthalpy entropy Gibbs free energy and calorimetry **Engineering Chemistry I (for BPUT)** Patrizia Bocchetta, 2024 *A Textbook of Physical Chemistry, 6th Edition* Sharma K.K. & Sharma L.K., 2016 A Textbook of Physical Chemistry *AP Chemistry Premium, 2024: 6 Practice Tests + Comprehensive Review + Online Practice* Neil D. Jespersen, Pamela Kerrigan, 2023-07-04 Always study with the most up to date prep Look for AP Chemistry Premium 2025 Prep Book with 6 Practice Tests Comprehensive Review Online Practice ISBN 9781506291802 on sale July 2 2024 Publisher's Note Products purchased from third party sellers are not guaranteed by the publisher for quality authenticity or access to any online entities included with the product *Chemistry for Degree Students B.Sc. (Honours) Semester II, 1/e (As per CBCS)* Madan R.L., 2022 This textbook has been designed to meet the needs of B Sc Honours Second Semester students of Chemistry as per the UGC Choice Based Credit System CBCS Maintaining the traditional approach to the subject this textbook lucidly explains the basics of Organic and Physical Chemistry Important topics such as alkanes alkenes alkynes stereochemistry aliphatic hydrocarbons thermochemistry chemical thermodynamics and chemical equilibrium are aptly discussed to give an overview of organic and physical chemistry Laboratory work has also been included to help students achieve solid conceptual understanding and learn experimental procedures **Engineering Chemistry** Shikha Agarwal, 2019-05-23 Gain a detailed understanding of the fundamental concepts of chemistry and their engineering applications with this fully revised second edition Catering to the needs of first and second semester undergraduate students from all branches of engineering taking courses on engineering chemistry it offers new material on topics such as periodic properties structure and bonding gaseous states ionic equilibrium oxidation and reduction Werner's coordination theory Sidgwick coordination theory valence bond theory crystal field theory bonding in coordination compounds and isomerism in

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