



NATIONAL
SENIOR CERTIFICATE

GRADE 11

NOVEMBER 2020

PHYSICAL SCIENCES P2
(CHEMISTRY)
EXEMPLAR

MARKS: 100
TIME: 1 hour



The question paper consists of 10 pages, including 4 extra sheets.
Use appropriate units for all calculations.

QUESTION 3

PHYSICAL SCIENCES PAPER 2

INTERMOLECULAR FORCES

GRADE 11

NOVEMBER 2020

Examiner's use: _____ Physical Sciences P2

QUESTION 3 starts on a NEW page.

3.1 The boiling points in the table below were obtained during an investigation carried out to measure the boiling points of pure 1-hexanol. The table also lists the ΔH_{vap} and ΔH_{fus} for 1-hexanol in kJ mol⁻¹ and in $^{\circ}\text{C}$ respectively.

Compound	Molecular mass (M _r)	Boiling point (T _b)
A	102	162
B	102	17

- 3.1.1 Define the term boiling point. (2)
- 3.1.2 Is pure 1-hexanol only composed of A and B? or A, B and C? or all of them? (2)
- 3.1.3 Name the type of intermolecular force that exists between molecules of both compounds A and B and its boiling point of these molecules. (2)
- 3.1.4 What kind of the compound A or B is? (1-hexanol) (pure) (1) (1-hexanol) (pure) (1) (1-hexanol) (pure) (1)
- 3.1.5 What compound B or C is? (1-hexanol) (pure) (1) (1-hexanol) (pure) (1) (1-hexanol) (pure) (1)
- 3.1.6 What compound B or C is? (1-hexanol) (pure) (1) (1-hexanol) (pure) (1) (1-hexanol) (pure) (1)

Compound C (M_r 102) is shown in the table. It has a smaller molecular mass than both compounds A and B but has a relatively higher boiling point of 162 $^{\circ}\text{C}$.

- 3.1.7 Explain why the boiling point of compound C is HIGHER than that of compound A and B by referring to the T_b and molecular mass of the compounds. (2)

3.2 Substances D and E are shown in the table below. At room temperature (25 $^{\circ}\text{C}$) and 1 atm, D is a solid and E is a liquid. The diagram below shows the structures of D and E.



- 3.2.1 Is substance D or E a solid, at the room temperature? (1)
- 3.2.2 Explain the reason for QUESTION 3.2.1 about the structure of the compounds. (2)

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Page 10 of 10

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Table of Contents Physical Science Paper2 November 2014

1. Understanding the eBook Physical Science Paper2 November 2014
 - The Rise of Digital Reading Physical Science Paper2 November 2014
 - Advantages of eBooks Over Traditional Books
2. Identifying Physical Science Paper2 November 2014
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Physical Science Paper2 November 2014
 - User-Friendly Interface
4. Exploring eBook Recommendations from Physical Science Paper2 November 2014
 - Personalized Recommendations
 - Physical Science Paper2 November 2014 User Reviews and Ratings
 - Physical Science Paper2 November 2014 and Bestseller Lists
5. Accessing Physical Science Paper2 November 2014 Free and Paid eBooks
 - Physical Science Paper2 November 2014 Public Domain eBooks
 - Physical Science Paper2 November 2014 eBook Subscription Services
 - Physical Science Paper2 November 2014 Budget-Friendly Options
6. Navigating Physical Science Paper2 November 2014 eBook Formats

- ePub, PDF, MOBI, and More
- Physical Science Paper2 November 2014 Compatibility with Devices
- Physical Science Paper2 November 2014 Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Physical Science Paper2 November 2014
 - Highlighting and Note-Taking Physical Science Paper2 November 2014
 - Interactive Elements Physical Science Paper2 November 2014
- 8. Staying Engaged with Physical Science Paper2 November 2014
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Physical Science Paper2 November 2014
- 9. Balancing eBooks and Physical Books Physical Science Paper2 November 2014
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Physical Science Paper2 November 2014
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Physical Science Paper2 November 2014
 - Setting Reading Goals Physical Science Paper2 November 2014
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Physical Science Paper2 November 2014
 - Fact-Checking eBook Content of Physical Science Paper2 November 2014
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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