

| Topic | Grade | What do you need to know | 😊 or 😊 | What do You have to work on? |
|----------------------|-------|--|--------|------------------------------|
| Atomic Structure | A | <ul style="list-style-type: none"> Why an atom is neutral in terms of sub-atomic particles Outline isotopes from data about numbers of protons and neutrons | | |
| | B | <ul style="list-style-type: none"> Know the charges and masses of protons, electrons and neutrons. Know how isotopes are the same but have different mass number. | | |
| | C | <ul style="list-style-type: none"> Describe the arrangement of an atoms electron structure - 2.8.8.2 Describe the arrangement of elements in the periodic table. | | |
| Ionic Bonding | A | <ul style="list-style-type: none"> Explain ionic bonding using the 'dot and cross' model. Explain the melting point of NaCl by the structure of and bonding. Explain why the melting point of NaCl is lower than that of MgO. | | |
| | B | <ul style="list-style-type: none"> Predict the properties of substances with giant ionic structures. Understand that atoms with a full outer shell are electronically stable. | | |
| | C | <ul style="list-style-type: none"> Ionic bonds are formed from metals and non-metals Electrons are donated and received. Metals always donate the electron(s). Non- metals always receive the electron(s) | | |
| The Periodic Table | A | <ul style="list-style-type: none"> Explain some covalent bonds using a 'dot and cross' model Explain double covalent binds using a 'dot and cross' model | | |
| | B | <ul style="list-style-type: none"> Know why CO₂ and H₂O have low melting points | | |
| | C | <ul style="list-style-type: none"> Covalent bonds always have non-metals and share electrons Know that the group number is the same as the number of electrons in the outer shell. Be able to find out how many shells an atom has from the period it is in. | | |
| The Group 1 Elements | A | <ul style="list-style-type: none"> Construct a balanced symbol equation for Cu²⁺, Fe²⁺ and Fe³⁺ to make Cu(OH)₂, Fe(OH)₂, Fe(OH)₃. Reactivity with water gives NaOH and KOH. Know the process is oxidation. | | |
| | B | <ul style="list-style-type: none"> Why group 1 elements form similar positive ions. Construct a balance symbol equation of group 1 with water | | |
| | C | <ul style="list-style-type: none"> Metal + Water → Metal Hydroxide + Hydrogen Know how to use a flame test to identify potassium, lithium and sodium compounds. | | |
| The Group 7 Elements | A | <ul style="list-style-type: none"> Predict if displacement reactions will occur with other halogens Explain the trend of reactivity of group 7 by ideas of electron gain Explain why a process is reduced from its ionic | | |

Ocr Gateway Separate Sciences B C4 Checklist

Camilla Rothe



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