

Chemical Equilibrium Lab

Introduction:

Some chemical systems reach the state of dynamic equilibrium involving interconverting reactants simultaneously forming as reactants or products and reverse. These reactions are called reversible reactions. In this reaction you observe an acid solution. Once the equilibrium is reached in the system the concentrations of reactants and products remain constant. The importance associated with equilibrium is a situation where the approach to a state state increases that reactions speed comes to a stop. Instead, the forward and reverse reactions continue to occur over time, what there is no net change in the concentrations of either the reactants or the products. Equilibrium conditions is reached only at equilibrium conditions or equilibrium can adjust the equilibrium. Le Chatelier's Principle states that applying stress changes in concentration or temperature or equilibrium will cause a shift in equilibrium for the change and allow the system to return to equilibrium. When some products are added an equilibrium reaction will shift to the reactants side to consume the added product and return to equilibrium. When more reactants are added, the equilibrium will shift to the products side to consume the added reactant. In this lab, you will be observing the changes of equilibrium when the amount of reactants, pH, temperature are changed. Follow Le Chatelier's Principle in the following reaction. Shifted equilibrium can be identified by the color change or solution according to the color indicators in the equations below. For example, if the color changes to red, the reaction shifts to the products side to consume the added reactant and is shifted to the products side.



Adding NaOH introduces OH^{-} ions that Fe^{3+} ions. This reaction Fe^{3+} reacts the present concentration shifts to reaction to consume the added OH}^{-} and is shifted to the products side.

Materials (Per Group):

- | | |
|---------------------------------------|---------------------------------------|
| 10 mL 0.1 M Fe^{3+} solution | 10 mL 0.1 M SCN^{-} solution |
| 10 mL 0.1 M NaOH solution | 10 mL 0.1 M HCl solution |
| 10 mL 0.1 M Fe^{3+} solution | 10 mL 0.1 M SCN^{-} solution |
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| 10 mL 0.1 M Fe^{3+} solution | 10 mL 0.1 M SCN^{-} solution |
| 10 mL 0.1 M NaOH solution | 10 mL 0.1 M HCl solution |

Observing Chemical Equilibrium Lab 46

K Morrison



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