

8.2 The Nature of Covalent Bonding

Connecting to Your World

You know that without oxygen to breathe, you could not live. But did you know that oxygen plays another important role in your life? High in the atmosphere, a different form of oxygen, called ozone, forms a layer that filters out harmful radiation from the sun. The colors in this map indicate the concentrations of ozone in various parts of Earth's atmosphere. In this section, you will learn how oxygen atoms can join in pairs to form the oxygen you breathe and can also join in groups of three oxygen atoms to form ozone.



The Octet Rule in Covalent Bonding

Recall that when ionic compounds form, electrons tend to be transferred so that each ion acquires a noble gas configuration. A similar rule applies for covalent bonds. **In covalent bonds, electron sharing usually occurs so that atoms attain the electron configurations of noble gases.** For example, each hydrogen atom has one electron. But a pair of hydrogen atoms share these two electrons when they form a covalent bond in a hydrogen molecule. Each hydrogen atom thus attains the electron configuration of helium, a noble gas with two electrons. Combinations of atoms of the nonmetals and metalloids in Groups 4A, 5A, 6A, and 7A of the periodic table are likely to form covalent bonds. In this case the atoms usually acquire a total of eight electrons, or an octet, by sharing electrons, so that the octet rule applies.

Single Covalent Bonds

The hydrogen atoms in a hydrogen molecule are held together mainly by the attraction of the shared electrons to the positive nuclei. Two atoms held together by sharing a pair of electrons are joined by a **single covalent bond**. Hydrogen gas consists of diatomic molecules whose atoms share only one pair of electrons, forming a single covalent bond.



Guide for Reading

Key Concepts

- What is the result of electron sharing in covalent bonds?
- How do electron-dot structures represent shared electrons?
- How do atoms form double or triple covalent bonds?
- How are coordinate covalent bonds different from other covalent bonds?
- How is the strength of a covalent bond related to its bond dissociation energy?
- How are oxygen atoms bonded in ozone?
- What are some exceptions to the octet rule?

Vocabulary

single covalent bond
structural formula
unshared pair
double covalent bond
triple covalent bond
coordinate covalent bond
polyatomic ion
bond dissociation energy
resonance structure

Reading Strategy

Identifying Main Idea/Details

List the main idea in the paragraph following the heading **The Octet Rule in Covalent Bonding**. As you read, list examples of how this rule is followed when a single covalent bond, a double covalent bond, and a triple covalent bond form.

Reading For 8 2 The Nature Of Covalent Bonding

Melvin Dresher



Reading For 8 2 The Nature Of Covalent Bonding:

Symmetry and Physical Properties of Crystals Cécile Malgrange, Christian Ricolleau, Michel Schlenker, 2014-12-04
Crystals are everywhere from natural crystals minerals through the semiconductors and magnetic materials in electronic devices and computers or piezoelectric resonators at the heart of our quartz watches to electro optical devices
Understanding them in depth is essential both for pure research and for their applications This book provides a clear thorough presentation of their symmetry both at the microscopic space group level and the macroscopic point group level The implications of the symmetry of crystals for their physical properties are then presented together with their mathematical description in terms of tensors The conditions on the symmetry of a crystal for a given property to exist then become clear as does the symmetry of the property The geometrical representation of tensor quantities or properties is presented and its use in determining important relationships emphasized An original feature of this book is that most chapters include exercises with complete solutions This allows readers to test and improve their understanding of the material The intended readership includes undergraduate and graduate students in materials science and materials related aspects of electrical and optical engineering researchers involved in the investigation of the physical properties of crystals and the design of applications based on crystal properties such as piezoelectricity electro optics optical activity and all those involved in the characterization of the structural properties of materials University of Michigan Official Publication ,1949

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