

Question number	Answer	Marks	Guidance
1	$\text{Ca} : \text{N} : \text{O} = 30.35/40.1 : 21.20/14.0 : 48.45/16.0$ $= 0.7569 : 1.5286 : 3.028$ Formula = CaN_2O_4	B1 B1	
2 (a)	$2\text{KClO}_3(\text{s}) \rightarrow 2\text{KCl}(\text{s}) + 3\text{O}_2(\text{g})$	B1	
2 (b)	$4.50 \times 10^{-3} \text{ mol}$	B1	
2 (c)	$n(\text{KClO}_3) = 3.0 \times 10^{-3} \text{ mol}$ $M(\text{KClO}_3) = 122.6 \text{ g mol}^{-1}$ mass of $\text{KClO}_3 = 0.3678 \text{ g}$	B1 B1 B1	
3	$M(\text{ZnSO}_4) = 161.5 \text{ g mol}^{-1}$ $n(\text{ZnSO}_4) = 6.57 \times 10^{-3} \text{ mol}$ $n(\text{H}_2\text{O}) = (1.893 - 1.061)/18 = 4.62 \times 10^{-2} \text{ mol}$ $x = n(\text{H}_2\text{O})/n(\text{ZnSO}_4) = 7$ Formula = $\text{ZnSO}_4 \cdot 7\text{H}_2\text{O}$	B1 B1 B1 B1	
4	$M(\text{Na}_2\text{CO}_3) = 106.0 \text{ g mol}^{-1}$ Actual $n(\text{Na}_2\text{CO}_3) = 0.0234 \text{ mol}$ Theoretical $n(\text{Na}_2\text{CO}_3) = 0.0234 \times 100/65$ $= 0.0360 \text{ mol}$ Theoretical $n(\text{NaHCO}_3) = 0.0720 \text{ mol}$ Mass of $\text{NaHCO}_3 = 0.0720 \times 84.0 = 6.05 \text{ g}$	B1 B1 B1 B1 B1	
5 (a)	$\text{Fe}_2\text{O}_3(\text{s}) + 3\text{CO}(\text{g}) \rightarrow 2\text{Fe}(\text{s}) + 3\text{CO}_2(\text{g})$	B1	
5 (b)	$M(\text{Fe}_2\text{O}_3) = 159.6 \text{ g mol}^{-1}$ $n(\text{Fe}_2\text{O}_3) = (10\,000 \times 10^6)/159.6 = 6.266 \times 10^7 \text{ mol}$ $n(\text{Fe}) = 1.253 \times 10^8 \text{ mol} = 6.992 \times 10^5 \text{ g}$ (6992 tonnes)	B1 B1 B1	
6 (a)	$\text{Na} : \text{N} : \text{O} = 27.1/23.0 : 16.5/14.0 : 56.4/16.0$ $= 1.178 : 1.179 : 3.525$ Formula = NaNO_3	B1 B1	
6 (b)	$2\text{NaNO}_3(\text{s}) \rightarrow 2\text{NaNO}_2(\text{s}) + \text{O}_2(\text{g})$	B1	

Ocr Chemistry Exam Question Answers

Mike Smith

A red circular graphic with a gradient, appearing as a semi-circle or a partial circle, located to the right of the 'Mike Smith' text bar.

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Revise AS Chemistry for Salters (OCR) Ann Daniels, 2005-02-22 Help students pull together the chemical ideas in the course and apply them to fresh contexts in exam questions **OCR(A) AS**

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level Chemistry Student Book 1 Mike Smith, John Older, 2015-06-26 Exam Board OCR Level A level Subject Chemistry First Teaching September 2015 First Exam June 2016 This is an OCR endorsed resource Stretch and challenge your students

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OCR A Chemistry AS Student Unit Guide: F321 New Edition Atoms, Bonds and Groups Mike Smith,2012-08-31 Written by a former senior examiner Mike Smith this OCR A AS Chemistry Student Unit Guide is the essential study companion for Unit F321 Atoms Bonds and Groups This full colour book includes all you need to know to prepare for your unit exam clear guidance on the content of the unit with topic summaries knowledge check questions and a quick reference index examiner s advice throughout so you will know what to expect in the exam and will be able to demonstrate the skills required exam style questions with graded student responses so you can see clearly what is required to get a better grade

A Level Chemistry For OCR Pasan Witharana,2022-06 Prepare for OCR A Level Chemistry with over 500 questions answers and mind maps covering various topics Learn everything you need to know for your exams while working on your exam technique This book aims to maximise your revision time by following the high yield revision approach Edited by a specialist OCR examiner the answers in the book follow the marking scheme answer format Learn the subject content through questions and answers Mind maps to consolidate learning for effective revision and maximise study time Practice all question styles including multiple choice extended answer and multi step calculations Learn to answer effectively to score the maximum marks Answers to all practise questions are provided Develop your exam technique and face the exam with confidence

OCR(B) A2 Chemistry (Salters) Student Unit Guide: Unit F335 Chemistry by Design Frank Harriss,2009-06-26 Student Unit Guides

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OCR AS Chemistry Student Unit Guide Mike Smith, 2009-06-26

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