

16. The dissociation of water is shown below:



- At 25 °C, the auto product of water, K_w , is 1.0×10^{-14} mol² dm⁻⁶.
At 65 °C, the auto product of water, K_w , is 1.0×10^{-13} mol² dm⁻⁶.

17. Explain what the dissociation of water is an indication of water's chemical nature.

Endothermic.

18. Predict, using a calculation, whether a solution of 1.0 mol dm⁻³ is neutral, acidic or alkaline.

Acidic

Explain by calculating pH.

19. pK_w , pH and pOH are significant notions.

Calculate pK_w at 65 °C.

Give your answer to two decimal places.

$$-\log(1.0 \times 10^{-13}) = 13.00$$

$$pK_w = 13.00$$

20. A solution of hydrochloric acid is mixed with water and the resulting solution is neutral at 25 °C.

Calculate the pH of the resulting solution at 65 °C.

Give your answer to two decimal places.

$$pH = 7.00$$

pH = _____ (1)

(Total: 10)

Ocr Chemistry A2 Unofficial Mark Scheme 2014

Daniel F McAuley



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Table of Contents Ocr Chemistry A2 Unofficial Mark Scheme 2014

1. Understanding the eBook Ocr Chemistry A2 Unofficial Mark Scheme 2014
 - The Rise of Digital Reading Ocr Chemistry A2 Unofficial Mark Scheme 2014
 - Advantages of eBooks Over Traditional Books
2. Identifying Ocr Chemistry A2 Unofficial Mark Scheme 2014
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Ocr Chemistry A2 Unofficial Mark Scheme 2014
 - User-Friendly Interface
4. Exploring eBook Recommendations from Ocr Chemistry A2 Unofficial Mark Scheme 2014
 - Personalized Recommendations
 - Ocr Chemistry A2 Unofficial Mark Scheme 2014 User Reviews and Ratings
 - Ocr Chemistry A2 Unofficial Mark Scheme 2014 and Bestseller Lists

5. Accessing Ocr Chemistry A2 Unofficial Mark Scheme 2014 Free and Paid eBooks
 - Ocr Chemistry A2 Unofficial Mark Scheme 2014 Public Domain eBooks
 - Ocr Chemistry A2 Unofficial Mark Scheme 2014 eBook Subscription Services
 - Ocr Chemistry A2 Unofficial Mark Scheme 2014 Budget-Friendly Options
6. Navigating Ocr Chemistry A2 Unofficial Mark Scheme 2014 eBook Formats
 - ePub, PDF, MOBI, and More
 - Ocr Chemistry A2 Unofficial Mark Scheme 2014 Compatibility with Devices
 - Ocr Chemistry A2 Unofficial Mark Scheme 2014 Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Ocr Chemistry A2 Unofficial Mark Scheme 2014
 - Highlighting and Note-Taking Ocr Chemistry A2 Unofficial Mark Scheme 2014
 - Interactive Elements Ocr Chemistry A2 Unofficial Mark Scheme 2014
8. Staying Engaged with Ocr Chemistry A2 Unofficial Mark Scheme 2014
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Ocr Chemistry A2 Unofficial Mark Scheme 2014
9. Balancing eBooks and Physical Books Ocr Chemistry A2 Unofficial Mark Scheme 2014
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Ocr Chemistry A2 Unofficial Mark Scheme 2014
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Ocr Chemistry A2 Unofficial Mark Scheme 2014
 - Setting Reading Goals Ocr Chemistry A2 Unofficial Mark Scheme 2014
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Ocr Chemistry A2 Unofficial Mark Scheme 2014
 - Fact-Checking eBook Content of Ocr Chemistry A2 Unofficial Mark Scheme 2014
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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