

Question	Answer	Mark	Comments
1	Planck's constant, $h = 6.63 \times 10^{-34} \text{ J s}$ Frequency, $f = 5.0 \times 10^{14} \text{ Hz}$ Energy, $E = hf$ $E = 3.315 \times 10^{-19} \text{ J}$	1 1 1 1	
2(a)	Electron has longest wavelength	1	1st mark: correct wavelength
2(b)	5.0×10^{14}	1	
2(c)	Planck's constant, $h = 6.63 \times 10^{-34} \text{ J s}$ Frequency, $f = 5.0 \times 10^{14} \text{ Hz}$ Energy, $E = hf$ $E = 3.315 \times 10^{-19} \text{ J}$	1 1 1 1	
2(d)	Electron has shortest wavelength	1	
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2(g)	Electron has shortest wavelength	1	
2(h)	Planck's constant, $h = 6.63 \times 10^{-34} \text{ J s}$ Frequency, $f = 5.0 \times 10^{14} \text{ Hz}$ Energy, $E = hf$ $E = 3.315 \times 10^{-19} \text{ J}$	1 1 1 1	
2(i)	Electron has shortest wavelength	1	
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2(y)	Electron has shortest wavelength	1	
2(z)	Planck's constant, $h = 6.63 \times 10^{-34} \text{ J s}$ Frequency, $f = 5.0 \times 10^{14} \text{ Hz}$ Energy, $E = hf$ $E = 3.315 \times 10^{-19} \text{ J}$	1 1 1 1	

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