

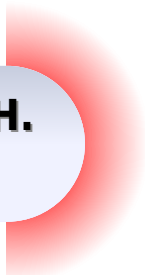
Chapter 1

Gases and the Zeroth Law of Thermodynamics

- 1.2.** A system is any part of the universe under observation. The “surroundings” includes everything else in the universe. Consider a solution calorimeter in which two aqueous solutions are mixed and temperature changes are recorded. In this case the solutes (reactants and products) would be considered the “system”. The water, calorimeter, and rest of the lab and world would be the “surroundings”.
- 1.4.** (a) $12.56 \text{ L} \times \frac{1000 \text{ mL}}{1 \text{ L}} \times \frac{1 \text{ cm}^3}{1 \text{ mL}} = 12,560 \text{ cm}^3 = 1.256 \times 10^4 \text{ cm}^3$
(b) $45^\circ\text{C} + 273.15 = 318 \text{ K}$
(c) $1.055 \text{ atm} \times \frac{1.01325 \text{ bar}}{1 \text{ atm}} \times \frac{100,000 \text{ Pa}}{1 \text{ bar}} = 1.069 \times 10^5 \text{ Pa}$
(d) $1233 \text{ mmHg} \times \frac{1 \text{ torr}}{1 \text{ mmHg}} \times \frac{1 \text{ atm}}{760 \text{ torr}} \times \frac{1.01325 \text{ bar}}{1 \text{ atm}} = 1.644 \text{ bar}$
(e) $125 \text{ mL} \times \frac{1 \text{ cm}^3}{1 \text{ mL}} = 125 \text{ cm}^3$
(f) $4.2 \text{ K} - 273.15 = -269.0^\circ\text{C}$
(g) $25750 \text{ Pa} \times \frac{1 \text{ bar}}{100,000 \text{ Pa}} = 0.2575 \text{ bar}$
- 1.6.** $p_{\text{tot}} = p_{\text{atmosph}} + \rho hg$, where ρhg correspond to the pressure exerted by the liquid
 $p_{\text{tot}} - p_{\text{atmosph}} = \rho hg = (1.0 \times 10^3 \text{ kg/m}^3)(0.23 \text{ m})(9.80 \text{ m/s}^2) = 2254 \text{ N/m}^2 = 2.3 \times 10^3 \text{ Pa}$
- 1.8.** In terms of the zeroth law of thermodynamics, heat will flow from the (hot) burner or flame on the stove into the (cold) water, which gets hotter. Then heat will move from the hot water into the (colder) egg.
- 1.10.** For this sample of gas under these conditions, $F(T) = 2.97 \text{ L} \times 0.0553 \text{ atm} = 0.164 \text{ L} \cdot \text{atm}$. If the pressure were increased to 1.00 atm: $0.164 \text{ L} \cdot \text{atm} = (1.00 \text{ atm}) \times V$; therefore $V = 0.164 \text{ L}$.
- 1.12.** $V = R \cdot \frac{nT}{p}$, which rearranges to $R = \frac{pV}{nT}$.
Therefore: $R = \frac{(2.66 \text{ bar})(27.5 \text{ L})}{(1.887 \text{ mol})(466.9 \text{ K})} = 0.0830 \frac{\text{L} \cdot \text{bar}}{\text{mol} \cdot \text{K}}$

Physical Chemistry David W Ball Solution Manual

**Hans Kuhn,H.D. Foersterling,David H.
Waldeck,Jeffry D. Madura**



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Student Solutions Manual, Physical Chemistry Second Edition, David W. Ball Jörg C. Woehl, 2015 *Physical Chemistry + Student Solutions Manual*, *Solutions Manual for Principles of Physical Chemistry, 3rd Edition* Hans Kuhn, David H. Waldeck, Horst-Dieter Försterling, 2024-10-29 This is a Solutions Manual to Accompany with solutions to the exercises in the main volume of Principles of Physical Chemistry Third Edition This book provides a unique approach to introduce undergraduate students to the concepts and methods of physical chemistry which are the foundational principles of Chemistry The book introduces the student to the principles underlying the essential sub fields of quantum mechanics atomic and molecular structure atomic and molecular spectroscopy statistical thermodynamics classical thermodynamics solutions and equilibria electrochemistry kinetics and reaction dynamics macromolecules and organized molecular assemblies Importantly the book develops and applies these principles to supramolecular assemblies and supramolecular machines with many examples from biology and nanoscience In this way the book helps the student to see the frontier of modern physical chemistry developments The book begins with a discussion of wave particle duality and proceeds systematically to more complex chemical systems in order to relate the story of physical chemistry in an intellectually coherent manner The topics are organized to correspond with those typically given in each of a two course semester sequence The first 13 chapters present quantum mechanics and spectroscopy to describe and predict the structure of matter atoms molecules and solids Chapters 14 to 29 present statistical thermodynamics and kinetics and applies their principles to understanding equilibria chemical transformations macromolecular properties and supramolecular machines Each chapter of the book begins with a simplified view of a topic and evolves to more rigorous description in order to provide the student and instructor flexibility to choose the level of rigor and detail that suits them best The textbook treats important new directions in physical chemistry research including chapters on macromolecules principles of interfaces and films for organizing matter and supramolecular machines as well as including discussions of modern nanoscience spectroscopy and reaction dynamics throughout the text

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