

Chapter 12: Atoms and the Periodic Table
Subatomic Particles

Three fundamental particles of the atom are the _____, _____, and _____. At the center of each atom lies the atomic _____, which consists of _____ and _____. The atomic number refers to the number of _____ in the nucleus. All atoms of the same element have the same number of _____, hence, the same atomic number.

Isotopes are atoms that have the same number of _____ but a different number of _____. An isotope is identified by its atomic mass number, which is the total number of _____ and _____ in the nucleus. A carbon isotope that has 6 _____ and 6 _____ is identified as carbon-12, where 12 is the atomic mass number. A carbon isotope having 6 _____ and 8 _____, on the other hand, is carbon-14.

1. Complete the following table:

Isotope	Number of...		
	Electrons	Protons	Neutrons
Hydrogen-1	1		
Chlorine-36		17	
Nitrogen-14			7
Potassium-40	19		
Arsenic-75		33	
Gold-197			118

2. Which results in a more valuable product — adding or subtracting protons from gold nuclei?

3. Which has more mass, a helium atom or a neon atom?

4. Which has a greater number of atoms, a gram of helium or a gram of neon?



Physical Science Chapter 12 Answers

Laurie E. Westphal



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Table of Contents Physical Science Chapter 12 Answers

1. Understanding the eBook Physical Science Chapter 12 Answers
 - The Rise of Digital Reading Physical Science Chapter 12 Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Physical Science Chapter 12 Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Physical Science Chapter 12 Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Physical Science Chapter 12 Answers
 - Personalized Recommendations
 - Physical Science Chapter 12 Answers User Reviews and Ratings
 - Physical Science Chapter 12 Answers and Bestseller Lists
5. Accessing Physical Science Chapter 12 Answers Free and Paid eBooks
 - Physical Science Chapter 12 Answers Public Domain eBooks
 - Physical Science Chapter 12 Answers eBook Subscription Services
 - Physical Science Chapter 12 Answers Budget-Friendly Options
6. Navigating Physical Science Chapter 12 Answers eBook Formats

- ePub, PDF, MOBI, and More
 - Physical Science Chapter 12 Answers Compatibility with Devices
 - Physical Science Chapter 12 Answers Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Physical Science Chapter 12 Answers
 - Highlighting and Note-Taking Physical Science Chapter 12 Answers
 - Interactive Elements Physical Science Chapter 12 Answers
 8. Staying Engaged with Physical Science Chapter 12 Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Physical Science Chapter 12 Answers
 9. Balancing eBooks and Physical Books Physical Science Chapter 12 Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Physical Science Chapter 12 Answers
 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
 11. Cultivating a Reading Routine Physical Science Chapter 12 Answers
 - Setting Reading Goals Physical Science Chapter 12 Answers
 - Carving Out Dedicated Reading Time
 12. Sourcing Reliable Information of Physical Science Chapter 12 Answers
 - Fact-Checking eBook Content of Physical Science Chapter 12 Answers
 - Distinguishing Credible Sources
 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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