

Study Guide: Chemistry Semester 1

ATOMIC & MOLECULAR STRUCTURE STANDARD 1

Atomic number

1. What is **atomic number**?
2. What is atomic number for the following elements?

Cr	Iron
Carbon	C

3. How does atomic number change moving left to right across the periodic table?

Atomic Weight (Atomic Mass)

4. What is the **atomic mass** for the following elements?

V	Oxygen
Beryllium	Mn

5. It was found that the periodic table was organized by increasing atomic number rather than mass. Give an example of a pair of elements where one element has a larger atomic number but a smaller atomic mass.
6. Identify the letters in the following periodic square:

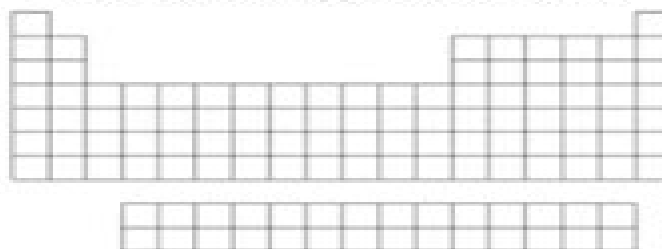
A →	16	
	S	← C
B →	32.066	
	Sulfur	← D

Atomic structure

1. Draw an atomic model of **Carbon-13**.
 In the drawing include the following:
 - Location of nucleus & electron cloud.
 - Location of 3 subatomic particles.
 - Amount of each subatomic particles.
 - A note where most of the atom's mass is found.

Organization of the Periodic Table

1. Identify on the periodic table below the location of
The 3 different types of elements.



2. Where are **groups** on the periodic table?
3. Where are **periods** on the periodic table?

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John Moore



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