

CHAPTER 5 REVIEW*The Periodic Law***SECTION 1****SHORT ANSWER** Answer the following questions in the space provided.

1. c In the modern periodic table, elements are ordered
(a) according to decreasing atomic mass.
(b) according to Mendeleev's original design.
(c) according to increasing atomic number.
(d) based on when they were discovered.
2. d Mendeleev noticed that certain similarities in the chemical properties of elements appeared at regular intervals when the elements were arranged in order of increasing
(a) density. (c) atomic number.
(b) reactivity. (d) atomic mass.
3. b The modern periodic law states that
(a) no two electrons with the same spin can be found in the same place in an atom.
(b) the physical and chemical properties of an element are functions of its atomic number.
(c) electrons exhibit properties of both particles and waves.
(d) the chemical properties of elements can be grouped according to periodicity, but physical properties cannot.
4. c The discovery of the noble gases changed Mendeleev's periodic table by adding a new
(a) period. (c) group.
(b) series. (d) level.
5. d The most distinctive property of the noble gases is that they are
(a) metallic. (c) metalloid.
(b) radioactive. (d) largely unreactive.
6. c Lithium, the first element in Group 1, has an atomic number of 3. The second element in this group has an atomic number of
(a) 4. (c) 11.
(b) 10. (d) 18.
7. An isotope of fluorine has a mass number of 19 and an atomic number of 9.
9 a. How many protons are in this atom?
10 b. How many neutrons are in this atom?
 ${}^{19}_{9}\text{F}$ c. What is the nuclear symbol of this fluorine atom, including its mass number and atomic number?

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GSNotes



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Table of Contents Periodic Law Section Quiz Answers

1. Understanding the eBook Periodic Law Section Quiz Answers
 - The Rise of Digital Reading Periodic Law Section Quiz Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Periodic Law Section Quiz Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Periodic Law Section Quiz Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Periodic Law Section Quiz Answers
 - Personalized Recommendations
 - Periodic Law Section Quiz Answers User Reviews and Ratings
 - Periodic Law Section Quiz Answers and Bestseller Lists
5. Accessing Periodic Law Section Quiz Answers Free and Paid eBooks
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 - Periodic Law Section Quiz Answers Compatibility with Devices
 - Periodic Law Section Quiz Answers Enhanced eBook Features
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 - Highlighting and Note-Taking Periodic Law Section Quiz Answers
 - Interactive Elements Periodic Law Section Quiz Answers
8. Staying Engaged with Periodic Law Section Quiz Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Periodic Law Section Quiz Answers
9. Balancing eBooks and Physical Books Periodic Law Section Quiz Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Periodic Law Section Quiz Answers
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Periodic Law Section Quiz Answers
 - Setting Reading Goals Periodic Law Section Quiz Answers
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Periodic Law Section Quiz Answers
 - Fact-Checking eBook Content of Periodic Law Section Quiz Answers
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements

- Interactive and Gamified eBooks

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