

Percent Yield Practice

3 pages + 3-page Answer KEY + Periodic Table

Chemistry

PERCENT YIELD

Show all work and circle or highlight your final answer.

$$\text{percent yield (\%)} = \frac{\text{actual yield}}{\text{theoretical yield}} \times 100$$

- 1 16.8 grams of sodium reacts with excess chlorine gas.



The theoretical yield of sodium chloride is 25.5 grams. The actual yield was 23.5 grams. Calculate the percent yield of sodium chloride.

- 2 4.8 grams of carbon reacts with excess oxygen gas.



The theoretical yield of carbon dioxide is 25.5 g. Calculate the percent yield of carbon dioxide.

- 3 1.6 grams of zinc reacts with excess sulfur.



The theoretical yield of zinc sulfide is 15.5 grams. Calculate the percent yield of zinc sulfide.

- 4 1.0 grams of phosphorus reacts with excess chlorine.



The theoretical yield of phosphorus trichloride is 4.5 grams. Calculate the percent yield of phosphorus trichloride.

Name _____

Date _____

- 5 Hydrogen gas (H_2) reacts with oxygen gas (O_2) to produce water (H_2O). 5.8 grams of hydrogen reacts with excess oxygen gas.



What is the theoretical yield?

- 6 The actual yield was 16.8 grams. Calculate the percent yield.

- 7 Magnesium (Mg) reacts with hydrochloric acid (HCl) to produce magnesium chloride (MgCl_2) and hydrogen gas (H_2). 1.0 gram of magnesium reacts with excess hydrochloric acid.



What is the theoretical yield?

- 8 The actual yield was 40.2 grams. Calculate the percent yield.

Name _____

Date _____

- 9 Calcium carbonate (CaCO_3) reacts with hydrochloric acid (HCl) to produce calcium chloride (CaCl_2), carbon dioxide (CO_2), and water (H_2O). 15.8 grams of calcium carbonate reacts with excess hydrochloric acid.



What is the theoretical yield of calcium chloride?

$$\begin{array}{r} 15.8 \text{ g CaCO}_3 \times \frac{1 \text{ mol CaCO}_3}{100.09 \text{ g CaCO}_3} \times \frac{1 \text{ mol CaCl}_2}{1 \text{ mol CaCO}_3} \times \frac{110.98 \text{ g CaCl}_2}{1 \text{ mol CaCl}_2} = 17.0 \text{ g CaCl}_2 \end{array}$$

$$\begin{array}{l} \text{Molar Mass CaCl}_2 \\ \text{Ca} = 40.08 \\ \text{Cl} = 35.45 \\ 40.08 + 2(35.45) = 110.98 \text{ g} \end{array}$$

10 The actual yield was 15.8 grams. Calculate the percent yield of calcium chloride.

$$\begin{array}{r} 15.8 \text{ g CaCl}_2 \times \frac{1 \text{ mol CaCl}_2}{110.98 \text{ g CaCl}_2} \times \frac{1 \text{ mol CaCO}_3}{1 \text{ mol CaCl}_2} \times \frac{100.09 \text{ g CaCO}_3}{1 \text{ mol CaCO}_3} = 14.2 \text{ g CaCO}_3 \end{array}$$

$$\begin{array}{l} \text{Molar Mass CaCO}_3 \\ \text{Ca} = 40.08 \\ \text{C} = 12.01 \\ \text{O} = 16.00 \times 3 = 48.00 \\ 40.08 + 12.01 + 48.00 = 100.09 \text{ g} \end{array}$$

- 11 The actual yield was 22.5 grams. Calculate the percent yield of calcium chloride.

$$\text{percent yield (\%)} = \frac{22.5 \text{ g}}{100.09 \text{ g}} \times 100 = 22.5\%$$

- 12 Potassium (K) reacts with iodine (I_2) to produce potassium iodide (KI). 1.8 grams of potassium reacts with excess iodine.



What is the theoretical yield?

$$\begin{array}{r} 1.8 \text{ g K} \times \frac{1 \text{ mol K}}{39.10 \text{ g K}} \times \frac{2 \text{ mol KI}}{2 \text{ mol K}} \times \frac{166.10 \text{ g KI}}{1 \text{ mol KI}} = 4.7 \text{ g KI} \end{array}$$

$$\begin{array}{l} \text{Molar Mass KI} \\ \text{K} = 39.10 \\ \text{I} = 126.90 \\ 39.10 + 126.90 = 166.10 \text{ g} \end{array}$$

YIELD

ANSWER KEY INCLUDED

Percent Yield Example Answer Key

K. S. Verma

A red circular graphic with a gradient, appearing as a semi-circle or a partial circle, located to the right of the author's name.

Percent Yield Example Answer Key:

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Climate-resilience policies and investments for Egypt's agriculture sector: Sustaining productivity and food security Perez, Nicostrato D., Kassim, Yumna, Ringler, Claudia, Thomas, Timothy S., ElDidi, Hagar, Breisinger, Clemens, 2021-09-24 The importance of a resilient agriculture sector in providing food security livelihoods and household income was highlighted in many countries by the recent pandemic as was the capacity of the sector to cushion the negative impacts of the subsequent economic slowdown This has been the case in Egypt where agriculture has been resilient to the health crisis in comparison with the service and industry sectors Breisinger et al 2020 However the sector s resiliency is

gradually being corroded by climate change with lasting harmful effects for agriculture and food systems

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