

Guide

Wednesday 12 October 18:54

Camp Woodward - Season 6 - Slice of

1225 | FUEL TV | 18:50 - 19:00



On this episode, the campers get ready for the Big Pizza Jam. Jacob goes to the Dine Tour event with his uncle, Frankie practices his tricks and Johnny tries his best at BBQ.



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$$\text{Cu} + 2\text{AgNO}_3 \rightarrow \text{Cu(NO}_3)_2 + 2\text{Ag}$$

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$$3\text{CO} + 7\text{H}_2 \rightarrow \text{C}_3\text{H}_8 + 3\text{H}_2\text{O}$$

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web start with 8 3 mole h₂s and multiply it with 2 mole h₂o over 2 mole h₂s from the equation

$$8.3 \text{ mole H}_2\text{S} \times \frac{2 \text{ mole H}_2\text{O}}{2 \text{ mole H}_2\text{S}} = 8.3 \text{ mole H}_2\text{O}$$

 8.3 mole h₂s x 18g h₂o i mole h₂o 149 4g h₂o is your theoretical yield so percent yield 137 1g h₂o actual yield 149 4g h₂o x 100
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$$\text{C}_3\text{H}_6 + 3\text{O}_2 \rightarrow 3\text{CO} + 3\text{H}_2\text{O}$$

 if the reaction of 21 3 grams of o₂ produces a 47 3 yield how many grams of co would be produced step one figure out how many moles of the limiting reagent you have

$$21.3 \text{ grams O}_2 \times \frac{1 \text{ mole O}_2}{32 \text{ grams O}_2} = 0.666 \text{ moles O}_2$$

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$$\text{Percent Yield} = \frac{\text{calculated mass of the product}}{\text{mass of the product expected}} \times 100$$

 in this reaction a yield 80 is anticipated

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80 g I_2O_5 1 mol I_2O_5 1 mol I_2 xs 1 333.8 g I_2O_5 1 mol I_2O_5 28 g CO 1 mol CO

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