

b) How many grams of phosphorus are required to react completely with 20.0 g of O_2 ?

$$20.0 \text{ g } O_2 \times \frac{1 \text{ mol } O_2}{31.998 \text{ g } O_2} \times \frac{4 \text{ mol P}}{5 \text{ mol } O_2} \times \frac{30.973 \text{ g P}}{1 \text{ mol P}} = 15.5 \text{ g P}$$

c) What is the theoretical yield in grams if you combine the amounts of reactants in part b?

$$20.0 \text{ g } O_2 \times \frac{1 \text{ mol } O_2}{31.998 \text{ g } O_2} \times \frac{2 \text{ mol } P_2O_5}{5 \text{ mol } O_2} \times \frac{141.941 \text{ g } P_2O_5}{1 \text{ mol } P_2O_5} = 35.5 \text{ g } P_2O_5$$

OR

$$15.5 \text{ g P} \times \frac{1 \text{ mol P}}{30.973 \text{ g P}} \times \frac{2 \text{ mol } P_2O_5}{4 \text{ mol P}} \times \frac{141.941 \text{ g } P_2O_5}{1 \text{ mol } P_2O_5} = 35.5 \text{ g } P_2O_5$$

5. A gaseous mixture containing 5.00 moles of H_2 and 7.00 moles of Br_2 reacts to form HBr .

a) Write a balanced chemical equation for this reaction.



b) Which reactant is limiting?

1 to 1 ratio
5.00 mol H_2 & 7.00 mol Br_2
 $\therefore H_2$ is limiting

c) What is the theoretical yield for this reaction in moles?

$$5.00 \text{ mol } H_2 \times \frac{2 \text{ mol } HBr}{1 \text{ mol } H_2} = 10.0 \text{ mol } HBr$$

d) What is the theoretical yield for this reaction in grams?

$$10.0 \text{ mol } HBr \times \frac{80.912 \text{ g } HBr}{1 \text{ mol } HBr} = 809 \text{ g } HBr$$

6. Sodium reacts with molecular hydrogen to form sodium hydride. A reaction mixture contains 10.0 g Na and 0.0235 g H_2 .

a) Write a balanced chemical equation for this reaction.



b) Which reactant is limiting? H_2 is limiting

$$10.00 \text{ g Na} \times \frac{1 \text{ mol Na}}{22.9898 \text{ g Na}} = 0.4350 \text{ mol Na} \quad 0.0235 \text{ g } H_2 \times \frac{1 \text{ mol } H_2}{2.0158 \text{ g } H_2} = 0.0117 \text{ mol } H_2$$
$$\frac{0.4350}{2} = 0.2175 \quad \frac{0.0117}{1} = 0.0117$$

c) What is the theoretical yield for this reaction in grams?

$$0.0117 \text{ mol } H_2 \times \frac{2 \text{ mol NaH}}{1 \text{ mol } H_2} \times \frac{23.979 \text{ g NaH}}{1 \text{ mol NaH}} = 0.562 \text{ g NaH}$$

d) When this reaction is actually performed, 0.428 g of NaH is recovered. What is the percent yield of the reaction?

$$\frac{0.428 \text{ g}}{0.562} \times 100 = 76.2\%$$

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RJ Shavelson



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