

This is like a distribution in math. $2(x+3) \rightarrow 2x+6$
 Everything inside of the parentheses is multiplied by two.

- b. What does the subscripted "2" outside the parentheses of the chemical formula tell you about the compound?

The subscript "2" outside of $\text{Ba}(\text{NO}_3)_2$ the parentheses indicate that there are two of the polyatomic ion nitrate (NO_3^-).

20. How many atoms of each element are in one formula unit of ammonium phosphate, $(\text{NH}_4)_3\text{PO}_4$?

nitrogen

3

hydrogen

12

phosphorus

1

oxygen

4

21. A student writes the chemical formula for the ionic compound calcium hydroxide as CaOH_2 .

- a. Write the chemical formula for each ion in the compound.

Calcium: Ca^{+2}

Hydroxide: OH^{-1}

- b. Why is the student's chemical formula for the compound calcium hydroxide wrong?

The subscript of "2" after hydrogen (H) indicate two hydrogens, not two hydroxide polyatomic ions. Parentheses should be placed around OH^{-1} and a subscript of "2" should be placed outside of the parentheses.

22. Many of the chemical formulas in Model 3 include parentheses. Which one of the following rules summarizes the appropriate use of parentheses in ternary ionic compounds? For the three rules that do not apply in all cases, show at least one counter example from the chemical formulas in Model 3.

Parentheses are used around any ion that is used more than once in a formula unit.

↳ Not true - $\text{Al}_2(\text{CO}_3)_3$ - Al has more than one atom.

Parentheses are used around any polyatomic ion.

↳ Not true - KNO_3 - NO_3^{-1} has no parentheses because there is only one.

Parentheses are used around any polyatomic ion used more than once in a formula unit.

Parentheses are only used around polyatomic anions used more than once in a formula unit.

↳ Not true - $(\text{NH}_4)_3\text{PO}_4$ - ammonium is a polyatomic cation. (NH_4^{+1})

23. Write chemical formulas for the following ternary ionic compounds.

- a. Calcium sulfate



- b. Copper(II) nitrate



- c. Lithium phosphate



- d. Potassium permanganate



- e. Aluminum sulfite



- f. Magnesium bicarbonate



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1. Understanding the eBook Polyatomic Ion Packet Answers
 - The Rise of Digital Reading Polyatomic Ion Packet Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Polyatomic Ion Packet Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Polyatomic Ion Packet Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Polyatomic Ion Packet Answers
 - Personalized Recommendations
 - Polyatomic Ion Packet Answers User Reviews and Ratings
 - Polyatomic Ion Packet Answers and Bestseller Lists
5. Accessing Polyatomic Ion Packet Answers Free and Paid eBooks
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 - Highlighting and Note-Taking Polyatomic Ion Packet Answers
 - Interactive Elements Polyatomic Ion Packet Answers
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 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Polyatomic Ion Packet Answers
 9. Balancing eBooks and Physical Books Polyatomic Ion Packet Answers
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 - Creating a Diverse Reading Collection Polyatomic Ion Packet Answers
 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
 11. Cultivating a Reading Routine Polyatomic Ion Packet Answers
 - Setting Reading Goals Polyatomic Ion Packet Answers
 - Carving Out Dedicated Reading Time
 12. Sourcing Reliable Information of Polyatomic Ion Packet Answers
 - Fact-Checking eBook Content of Polyatomic Ion Packet Answers
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 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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