#### Chemistry -- Gas Laws ACT Questions [USE TO STUDY FOR YOUR TEST]

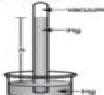
#### Multiple Choice

Identify the choice that best completes the statement or answers the question.

At 22°C, the air pressure in a car tire is 1293 torr.

Convert this pressure to atmospheres (atm).

- a. 12.76
- b. 5.262
- c. 1.701
- d. 1.293
- 2. This diagram shows a closed-end mercury (Hg)



barometer.

When atmospheric pressure is 1.30 atm, what is the height (h), in millimeters, of the Hg in the barometer?

- a. 132
- b. 585
- c. 760
- d. 988
- A chemist compresses a 1.5 L sample of carbon dioxide (CO<sub>2</sub>) gas at 1.2 atm to a final volume of 0.75 L. Assuming the temperature is constant, what is the final pressure, in atm, exerted by the CO<sub>2</sub> gas?
  - a. 0.60
  - b. 0.94
  - c. 1.7
  - d. 2.4
- A vessel connected to an open-end mercury (Hg)
  manometer contains nitrogen (N<sub>2</sub>) gas. The
  atmospheric pressure is 0.993 atm. The height
  difference between the two arms of the manometer



is 122 mm.

What is the pressure, in mm Hg, of the N2 gas?

- a. 355
- b. 633

- c. 755
- d. 877
- At 1.00 atm, a scaled weather balloon contains 20.0 L of helium (He) gas at 25.0°C. Assume that none of the He escapes and the pressure is constant. What is the volume, in liters, of He in the weather balloon at 35.0°C?
  - a. 40.0
  - b. 28.0
  - c. 20.7
  - d. 14.3
- 6. At 1254 m above sea level, the city of Helena, Montana, has an average atmospheric pressure of 0.859 atm. A chemistry student in Helena finds that pure distilled water boils at 95.9°C. Which statement explains the student's results at this high elevation?
  - The boiling point of water increases due to lower atmospheric pressure.
  - The boiling point of water increases due to higher atmospheric pressure.
  - The boiling point of water decreases due to lower atmospheric pressure.
  - The boiling point of water decreases due to higher atmospheric pressure.
- For a chemistry homework assignment, Beth must consider the effusion rates of methane (CH<sub>4</sub>) gas and nitrogen trifluoride (NF<sub>2</sub>) gas at 25°C. Effusion occurs as gas particles escape from a container through a small hole. Which statement accurately compares the effusion rates of CH<sub>4</sub> and NF<sub>3</sub> at 25°C?
  - a. CH<sub>4</sub> effuses 2 times faster than NF<sub>3</sub>.
  - b. CH<sub>4</sub> effuses 4 times faster than NF<sub>3</sub>.
  - c. NF3 effuses 2 times faster than CH4.
  - d. NF<sub>3</sub> effuses 4 times faster than CH<sub>4</sub>.
- At standard temperature and pessure, a chemist burns 3.50 L of pentane (C<sub>2</sub>H<sub>12</sub>) in excess oxygen (O<sub>2</sub>) gas. C<sub>2</sub>H<sub>12</sub>(I) + 8 O<sub>2</sub>(g) ⇒ 5 CO<sub>2</sub>(g) + 6 H<sub>2</sub>O(I) What is the maximum number of liters of carbon dioxide (CO<sub>2</sub>) that the reaction can produce?
  - a. 17.5
  - b. 8.50
  - c. 5.60
  - d. 0.700

# **Sample Test Questions On Gas Laws**

**The Princeton Review** 

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## **Table of Contents Sample Test Questions On Gas Laws**

- 1. Understanding the eBook Sample Test Questions On Gas Laws
  - The Rise of Digital Reading Sample Test Questions On Gas Laws
  - Advantages of eBooks Over Traditional Books
- 2. Identifying Sample Test Questions On Gas Laws
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
- 3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Sample Test Questions On Gas Laws
  - User-Friendly Interface
- 4. Exploring eBook Recommendations from Sample Test Questions On Gas Laws
  - Personalized Recommendations
  - Sample Test Questions On Gas Laws User Reviews and Ratings
  - Sample Test Questions On Gas Laws and Bestseller Lists

- 5. Accessing Sample Test Questions On Gas Laws Free and Paid eBooks
  - Sample Test Questions On Gas Laws Public Domain eBooks
  - Sample Test Questions On Gas Laws eBook Subscription Services
  - Sample Test Questions On Gas Laws Budget-Friendly Options
- 6. Navigating Sample Test Questions On Gas Laws eBook Formats
  - ∘ ePub, PDF, MOBI, and More
  - Sample Test Questions On Gas Laws Compatibility with Devices
  - Sample Test Questions On Gas Laws Enhanced eBook Features
- 7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Sample Test Questions On Gas Laws
  - Highlighting and Note-Taking Sample Test Questions On Gas Laws
  - Interactive Elements Sample Test Questions On Gas Laws
- 8. Staying Engaged with Sample Test Questions On Gas Laws
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - $\circ\,$  Following Authors and Publishers Sample Test Questions On Gas Laws
- 9. Balancing eBooks and Physical Books Sample Test Questions On Gas Laws
  - Benefits of a Digital Library
  - $\circ$  Creating a Diverse Reading Collection Sample Test Questions On Gas Laws
- 10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
- 11. Cultivating a Reading Routine Sample Test Questions On Gas Laws
  - Setting Reading Goals Sample Test Questions On Gas Laws
  - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Sample Test Questions On Gas Laws
  - Fact-Checking eBook Content of Sample Test Questions On Gas Laws
  - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
- Exploring Educational eBooks
- 14. Embracing eBook Trends
  - Integration of Multimedia Elements
  - Interactive and Gamified eBooks

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