

## Polyatomic Ions WS

Ammonium (NH<sub>4</sub>)<sup>1+</sup>  
Carbonate (CO<sub>3</sub>)<sup>2-</sup>  
Chromate (CrO<sub>4</sub>)<sup>2-</sup>

Dichromate (Cr<sub>2</sub>O<sub>7</sub>)<sup>2-</sup>  
Hydroxide (OH)<sup>1-</sup>  
Nitrate (NO<sub>3</sub>)<sup>1-</sup>

Nitrite (NO<sub>2</sub>)<sup>1-</sup>  
Sulfate (SO<sub>4</sub>)<sup>2-</sup>  
Sulfite (SO<sub>3</sub>)<sup>2-</sup>

Phosphate (PO<sub>4</sub>)<sup>3-</sup>

**Write the NAME of each of the following compounds.**

1. (NH<sub>4</sub>)Cl Ammonium Chloride
2. Be(SO<sub>4</sub>) Beryllium Sulfate
3. (NH<sub>4</sub>)<sub>3</sub>N Ammonium Nitride
4. MgCl<sub>2</sub> Magnesium Chloride
5. NH<sub>4</sub>(NO<sub>3</sub>) Ammonium Nitrate
6. Sr<sub>3</sub>(PO<sub>4</sub>)<sub>2</sub> Strontium Phosphate
7. Zn(CrO<sub>4</sub>)<sub>2</sub> Zinc Dichromate
8. K<sub>2</sub>(Cr<sub>2</sub>O<sub>7</sub>) Potassium Dichromate

9. Ga(ClO<sub>3</sub>)<sub>3</sub> Gallium Chlorate
10. Cu(OH) Copper I Hydroxide
11. (NH<sub>4</sub>)<sub>3</sub>(PO<sub>4</sub>) Ammonium Phosphate
12. Fe(SO<sub>4</sub>) Iron II Sulfate
13. Mg(NO<sub>2</sub>)<sub>2</sub> Magnesium Nitrite
14. (NH<sub>4</sub>)NO<sub>2</sub> Ammonium Nitrite
15. Na<sub>2</sub>(Cr<sub>2</sub>O<sub>7</sub>) Sodium Dichromate
16. Na(OH) Sodium Hydroxide

**Write the CHEMICAL FORMULA for each of the given NAMES.....cross charges**

- |                         |  |   |   |                     |
|-------------------------|--|---|---|---------------------|
| 17. calcium carbonate   | Ca <sup>2+</sup> (CO <sub>3</sub> ) <sup>2-</sup>              | = | Ca <sub>1</sub> (CO <sub>3</sub> ) <sub>1</sub> | = CaCO <sub>3</sub> |
| 18. barium nitrate      | Ba <sup>2+</sup> +2 NO <sub>3</sub> <sup>1-</sup>              | = | Ba(NO <sub>3</sub> ) <sub>2</sub>               |                     |
| 19. ammonium sulfate    | NH <sub>4</sub> <sup>1+</sup> +1 SO <sub>4</sub> <sup>2-</sup> | = | (NH <sub>4</sub> ) <sub>2</sub> SO <sub>4</sub> |                     |
| 20. aluminum hydroxide  | Al <sup>3+</sup> +3 OH <sup>1-</sup>                           | = | Al(OH) <sub>3</sub>                             |                     |
| 21. calcium phosphate   | Ca <sup>2+</sup> +2 PO <sub>4</sub> <sup>3-</sup>              | = | Ca <sub>3</sub> (PO <sub>4</sub> ) <sub>2</sub> |                     |
| 22. cesium nitrate      | Cs NO <sub>2</sub> <sup>1-</sup>                               | = | CsNO <sub>2</sub>                               |                     |
| 23. sodium nitrite      | Na <sup>1+</sup> +1 NO <sub>2</sub> <sup>1-</sup>              | = | NaNO <sub>2</sub>                               |                     |
| 24. calcium sulfate     | Ca <sup>2+</sup> +2 SO <sub>4</sub> <sup>2-</sup>              | = | CaSO <sub>4</sub>                               |                     |
| 25. beryllium sulfate   | Be SO <sub>4</sub> <sup>2-</sup>                               | = | BeSO <sub>4</sub>                               |                     |
| 26. sodium carbonate    | Na <sup>1+</sup> +1 CO <sub>3</sub> <sup>2-</sup>              | = | Na <sub>2</sub> CO <sub>3</sub>                 |                     |
| 27. magnesium phosphate | Mg <sup>2+</sup> +2 PO <sub>4</sub> <sup>3-</sup>              | = | Mg <sub>3</sub> (PO <sub>4</sub> ) <sub>2</sub> |                     |
| 28. calcium phosphate   | Ca <sup>2+</sup> +2 PO <sub>4</sub> <sup>3-</sup>              | = | Ca <sub>3</sub> (PO <sub>4</sub> ) <sub>2</sub> |                     |

# Polyatomic Ions Ws

**S Ashworth**

A red circular graphic with a gradient, appearing as a stylized arrow or a partial circle, located to the right of the author's name.

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